SCHEME & SYLLABUS

M.Sc (Hons.) Chemistry

(CHOICE BASED CREDIT SYSTEM)

Department of Physical Sciences

University Institute of Sciences and Humanities (UISH)

Sant Baba Bhag Singh University

2020

ABOUT THE DEPARTMENT

The Physical Sciences expands our knowledge of the universe and underlines new technologies, which benefit our society. In keeping with the heritage of imparting quality education, teaching and research are the prime motive of the Department of Physical Sciences.

Department of Physical Sciences is dynamic and progressive in its development of new course initiatives. The faculty is well placed to contribute substantially to the goal of SBBSU and becoming a research oriented organization. The teaching is by way of interactive sessions between students and teachers. Our courses ensure a coherent degree structure while encouraging interdisciplinary approach.

SALIENT FEATURES OF THE DEPARTMENT

- The Department has highly qualified, young, dynamic and dedicated Faculty in various fields of Physical Sciences *viz.* Chemistry, Physics, Mathematics.
- The Department has well equipped laboratories with a number of instruments and facilities like, UV- Visible Spectrophotometer, High Speed Centrifuge, Muffle furnance, Digital water bath, Polarimeter , Turbiditimeter, AbbsRefractrometer, Digital weighing balance, Magnetic plate with stirrer, pH meter, Conductometer, Flame Photometer, colorimeter and a double distillation plant, Spring balance , Sodium Lamp Transformer **,** Young's modulus,Ultrasonic interferometer **,** Rheostat **,** Maxwell needle apparatus kit **,** Magnetic field of solenoid **,** Ballistic Galvanometer **,** Deflection and vibration Magnetometer **,** Electron spin resonance
- The Department keeps its students abreast of latest advancements in technology through ultra-modern computer facilities, e-learning, virtual labs, SWAYAM Courses as per UGC guidelines.
- The department updates curricula on a regular basis to ensure that the students keep up with the changing trends of education and research globally. The syllabi of courses are designed to equip students to qualify exams such as GATE, UGC- NET / SLET, etc.
- Student centric, ICT enabled and interactive teaching
- Students and teachers participation in International, National, State and Regional seminars and conferences.
- Curricular and the co-curricular activities are well balanced in the Teaching Learning environment to provide holistic education to the students.
- Flexiblility in course curriculum as per the needs of students & PG Programmes with Project as research component.
- The outcome based teaching model of faculty comprising of theoretical work, regular academic activities such as research projects, seminars, resource learning and hands-on laboratory work.
- Along with Industry aligned academia, expert interaction, is the key features of the department.

***M.Sc (Hons.) Chemistry* :**

M.Sc (Hons.) Chemistry is a route for the science students who have completed their bachelor degree taking chemistry as one subjects to join ***M.Sc (Hons.) Chemistry program***. The program is designed to develop the knowledge of students about the field of chemical sciences and to served the society. They can also pursue Ph.D program in Chemistry after master degree.

Vision

To aspire, achieve and sustain for excellence in academics and research through scientific knowledge so as to provide solutions to global environmental issues and transform graduates into responsible citizens and competent professionals.

Mission:

- 1.Holistic development of learner through academic excellence, employability, acquisition of analytical skills and higher research.
- 2.To explore and advance new frontiers in physical sciences and integration with interdisciplinary sciences through visionary research for the benefit of society
- 3.To develop graduates for lifelong learning and professional growth.

ELIGIBILITY CRITERIA : B.Sc. with Chemistry as main subject

DURATION: 2 Years

CAREER PATHWAYS

The program is designed to meet the growing requirement of qualified professionals in field of chemical and pharmaceuticals industry and education. Master degree holders are hired both by Government and private organizations. They can join as Ph.D scholar further.

• Government Jobs

Prepare students for various government jobs such as education sector, Chemical industry, pharmaceuticals sector and civil services etc..

• Corporate Jobs

Multiple pathways designed according to the level of the students to prepare them for different job profiles as per needs of industrial sector.

• Higher Studies

This pathway prepares students for Higher Studies and helps in their research also.

• Entrepreneurship : To set up new ventures as analytical & Testing labs

Programme Educational Objective (PEO)

PEO1. To impart quality education in chemical sciences to achieve excellence in teaching-learning and research.

PEO2. To provide hand on training and execution of the chemical experiments and safe handling of chemistry laboratory and chemical waste.

PEO3. To construct a bridge between the theoretical and practical aspects of chemistry and inculcate research apptitude.

PEO4. To equip the learners to apply knowledge of Chemistry and to analyze the local and global impact of chemistry on individuals, organizations, and society.

PEO5. To develop talented and committed human resource which act as catalyst to support interdisciplinary research and become fit for industry and entrepreneur.

PEO6. To develop employable skills and life time leaning.

Programme Outcomes (PO)

PO1. Able to identify, analyze, think critically, formulate and solve chemical problems.

PO2. Efficiently work safely and execute chemical experiments using appropriate technologies/instrumentation for analysis and data interpretation.

PO3. Eligible for competetive examinations such as NET,SLET, GATE, for research and employment in Government and private organizations.

PO4. Able to function in multidisciplinary teams become equiped for industrial outreach and to meet the desired needs of society

PO5. Develop skills for critically review scientific information and become able to comprehend and write effective reports and design documentation.

PO6. Become Proficient and ethically responsible citizen.

Programme Specific Outcomes (PSO)

PSO1. Able to provide chemical nomenclature, classification, structure, reactivity and stereochemistry of organic and inorganic matter.

PSO2. Proficient in organic and inorganic reaction mechanisms and chemical analysis through quantitative/qualitative mode.

PSO3. Equiped and apply modern spectroscopic methods of analysis for chemical characterization od any form of matter.

PSO4. Employ core analytical and practical experiences of Chemical Sciences for the Societal expectations and solutions for environmental problems.

PSO5. Proficient in theoretical as well as practical aspects of Electrochemistry,chemical themodynamics, kinetics, quantum chemistry.

PSO6. Acquire ability to explain applications of Chemistry relates to the real world in term of advanced synthetic methods, advanced materials and analytical tools.

ABOUT THE CHOICE BASED CREDIT SYSTEM (CBCS)

The CBCS provides an opportunity for the students to choose courses from the prescribed courses comprising core, elective/minor or skill based courses. The courses can be evaluated following the grading system, which is considered to be better than the conventional marks system. The basic idea is to look into the needs of the students so as to keep up-to-date with development of higher education in India and abroad. CBCS aims to redefine the curriculum keeping pace with the liberalization and globalization in education. CBCS allows students an easy mode of mobility to various educational institutions spread across the world along with the facility of transfer of credits earned by students.

1. Curriculum Structure: *M.Sc (Hons) Chemistry* programme will have a curriculum with Syllabi consisting of following type of courses:

I. Ability Enhancement Courses (AEC): The Ability Enhancement Courses (AEC) may be of two kinds: Ability Enhancement Compulsory Courses (AECC) and Skill Enhancement Courses (SEC). AECC courses are the courses based upon the content that leads to Knowledge enhancement; these are mandatory for all disciplines.
SEC courses are value-based and/or skill-based and are aimed at providing hands-on-training, competencies, skills, etc.
A. Ability Enhancement Compulsory Courses (AECC): Environmental Science, English Communication/MIL Communication.

B. Skill Enhancement Courses (SEC): These courses may be chosen from a pool of courses designed to provide value-based and/or skill-based knowledge.

II. Core Courses (CR): A course, which should compulsorily by studied by a candidate as a core requirement is termed as a Core course. These courses are employability enhancement courses relevant to the chosen program of study. Program core comprises of Theory, Practical, Project, Seminar etc. Project work is considered as a special course involving application of knowledge in solving/ analyzing/exploring a real life situation/ difficult problem.

III. Elective Courses: Elective course is generally a course which can be chosen from a pool of courses and which may be very specific or specialized or advanced or supportive to the discipline/subject of study or with provides an extended scope or which enables an exposure to some other discipline/subject/domain or nurtures the candidate's proficiency/skill. Accordingly, elective course may be categorizes as:

A. Discipline Specific Elective (DSE) Course: Elective courses may be offered by the main discipline/subject of study is referred to as Discipline Specific Elective.

B. Project (I): An elective course designed to acquire special/advanced knowledge, such as supplement study/support study to a project work, and a candidate studies such a course on his own with an advisory support by a teacher/faculty member is called dissertation/project.

2. NOMENCLATURE USED:

CR: Core Course
AEC: Ability Enhancement Core Course
SEC: Skill Enhancement Core Course
EC: Elective Course

Index

S.No	Subject Type	Subject name	Subject Code	Semester	Page number
1.		**Scheme**		**1-VI**	
2.	CR	CHM501	Group Theory, Symmetry and Ligand Field Theory	**I**	
3.	CR	CHM503	Organic Reaction Mechanism-I	**I**	
4.	CR	CHM505	Thermodynamics: Chemical and Statistical Thermodynamics	**I**	
5.	CR	CHM507	Pericyclic reactions & Photochemistry (organic and Inorganic)	**I**	
6.	AECC	CSE551	Computers for Chemists-Theory	**I**	
7.	CR	CHM509	Inorganic Chemistry Practical-I	**I**	
8.	CR	CHM511	Organic Chemistry Practical-I	**I**	
9.	AECC	CSE553	Computers in Chemistry-Lab-I	**I**	
10.	CR	CHM502	Organometallics Chemistry and Metal Clusters	**II**	
11.	CR	CHM504	Organic Reaction Mechanism-II	**II**	
12.	CR	CHM506	Quantum Chemistry and Chemical Dynamics	**II**	
13.	CR	CHM508	Spectroscopy:1(Techniques for Structural elucidation of Organic Compounds)	**II**	
14.	SEC	MAT 528	Mathematics for Chemists(for B.Sc. Medical students)	**II**	
15.	CR	CHM528	Chemistry of biological systems (for B.Sc. Non Medical students)	**II**	
16.	SEC	CHM500	Seminar	**II**	
17.	CR	CHM510	Organic Chemistry Practical-II	**II**	
18.	CR	CHM512	Physical Chemistry Practical-II	**II**	
19.	SEC	PHY540	Research Methodology & Intellectual properties rights	**II**	
20.	CR	CHM601	Spectroscopy -2 (Techniques for Structural elucidation of Physical & Inorganic Compounds)	**III**	
21.	CR	CHM603	Electrochemistry & Surface Chemistry	**III**	
22.	EC	CHM*	**ELECTIVE I** (Choose any one) A. CHM605 (Advanced coordination Chemistry)	**III**	

			B. CHM607 (Environmental Chemistry) C. CHM609(Recent Trends in Inorganic Chemistry)		
23.	EC	CHM*	**ELECTIVE-II** Choose any one) A. CHM615 (Pharmaceutical Chemistry & Drug Design) B. CHM617 (Bio-Organic Chemistry) C. CHM619 (Material Chemistry)	**III**	
24.	SEC	CHM629	Project Part-I (Review of Litrature)	**III**	
25.	CR	CHM611	Inorganic Chemistry Practical-III	**III**	
26.	CR	CHM613	Physical Chemistry Practical-III	**III**	
27.	CR	CHM602	Chemistry of Natural Products/Heterocyclic Chemistry	**IV**	
28.	CR	CHM604	Bio-Inorganic Chemistry	**IV**	
29.	EC	CHM*	**ELECTIVE III** (Choose any one) A. CHM606 (Instrumental Methods of Analysis) B. CHM608 (Nano-Science& Nano Chemistry) C. CHM610(Green Chemistry) D. CHM618 (Photophysical Chemistry)	**IV**	
30.	EC	CHM*	**ELECTIVE-IV** Choose any one) A. CHM612 (Industrial Chemical analysis & Quality Control) B. CHM614 (Polymer Science) C. CHM616 (Chemistry of Materials) D. CHM620 (Biofuels)	**IV**	
31.	SEC	CHM630	Project Part-II	**IV**	
32.	AECC	BOT001	Natural Hazards and Disaster Management	**IV**	

CR: Core Course
EC: Elective Course
AECC: Ability Enhancement Core Course
SEC: Skill Enhancement Core Course

COURSE CLASSIFICATION

1.Core Courses (Semester I-IV)

	Subject name	Subject Code	Semester	Page number
1	CHM501	Group Theory, Symmetry and Ligand Field Theory	I	
2	CHM503	Organic Reaction Mechanism-I	I	
3	CHM505	Thermodynamics: Chemical and Statistical Thermodynamics	I	
4	CHM507	Pericyclic reactions & Photochemistry (organic and Inorganic)	I	
5	CHM509	Inorganic Chemistry Practical-1	I	
6	CHM511	Organic Chemistry Practical-1	I	
7	CHM502	Organometallics Chemistry and Metal Clusters	II	
8	CHM504	Organic Reaction Mechanism-II	II	
9	CHM506	Quantum Chemistry and Chemical Dynamics	II	
10	CHM508	Spectroscopy:1(Techniques for Structural elucidation of Organic Compounds)	II	
11	PHY 540	Research Methodology &Intellectual Properties Rights	II	
12	CHM510	Organic Chemistry Practical-II	II	
13	CHM512	Physical Chemistry Practical-I	II	
14	CHM601	Spectroscopy -2 (Techniques for Structural elucidation of Physical & Inorganic Compounds)	III	
15	CHM603	Electrochemistry & Surface Chemistry	III	
16	CHM611	Inorganic Chemistry Practical-II	III	
17	CHM613	Physical Chemistry Practical-II	III	
18	CHM602	Chemistry of Natural Products/Heterocyclic Chemistry	IV	
19	CHM604	Bio-Inorganic Chemistry	IV	

2. Discipline Elective Courses (Semester III-IV)

S.No	Subject name	Subject Code	Semester	Page number
1	CHM*	**ELECTIVE I** (Choose any one) D. CHM605 (Advanced coordination Chemistry) E. CHM607 (Environmental Chemistry) F. CHM609(Recent Trends in Inorganic Chemistry)	III	
2	CHM*	**ELECTIVE-II** Choose any one) D. CHM615 (Pharmaceutical Chemistry & Drug Design) E. CHM617 (Bio-Organic Chemistry) F. CHM619 (Material Chemistry)	III	
3	CHM*	**ELECTIVE III** (Choose any one) E. CHM606 (Instrumental Methods of Analysis) F. CHM608 (Nano-Science& Nano Chemistry) G. CHM610(Green Chemistry) H. CHM618 (Photophysical Chemistry)	IV	
4	CHM*	**ELECTIVE-IV** Choose any one) E. CHM612 (Industrial Chemical analysis & Quality Control) F. CHM614 (Polymer Science) G. CHM616 (Chemistry of Materials) H. CHM620 (Biofuels)	IV	

3. Skill enhancement courses (Semester I-IV)

S.No	Subject name	Subject Code	Semester	Page number
1	PHY 540	Research Methodology &Intellectual Properties Rights	II	
2	CHM500	Seminar	II	
3	CHM629	Project Part-I	III	
4	CHM630	Project Part-II	IV	

4. Ability Enhancement Courses (Semester 1-IV)

S.No	Subject name	Subject Code	Semester	Page number
1	CSE551	Computers for Chemists-Theory	I	
2	CSE553	Computers in Chemistry-Lab-1	I	
3	MAT528 / CHM528	Mathematics for Chemists(for B.Sc. Medical students)/ Chemistry of biological systems (for B.Sc. Non Medical students)	II	
4	BOT001	Natural Hazards and Disaster Management	IV	

Course Scheme (M. Sc. (Hons.) Chemistry)

Semester-I

I. Theory Subjects

Sr. No.	Type of course	Sub Code	Subject Name	Contact Hours (L:T:P)	Credits (L:T:P)	Total Contact Hours	Total Credit Hours
1	CR	CHM501	Group Theory, Symmetry and Ligand Field Theory	4:0:0	4:0:0	4	4
2	CR	CHM503	Organic Reaction Mechanism-I	4:0:0	4:0:0	4	4
3	CR	CHM505	Thermodynamics: Chemical and Statistical Thermodynamics	4:0:0	4:0:0	4	4
4	CR	CHM507	Pericyclic reactions & Photochemistry (organic and Inorganic)	4:0:0	4:0:0	4	4
5	AEC	CSE551	Computers for Chemists-Theory	3:0:0	3:0:0	3	3
		II. Practical Subjects					
6	CR	CHM509	Inorganic Chemistry Practical-I	0:0:4	0:0:4	4	2
7	CR	CHM511	Organic Chemistry Practical-I	0:0:4	0:0:4	4	2
8	AEC	CSE553	Computers in Chemistry-Lab-I	0:0:2	0:0:1	2	1
		Total				29	24

Total Credit Hours-24
Total Contact Hours- 29

CR: Core Course
AEC: Ability Enhancement Core Course
SEC: Skill Enhancement Core Course

Semester-II

I. Theory Subjects

Sr. No.	Type of course	Sub Code	Subject Name	Contact Hours (L:T:P)	Credits (L:T:P)	Total Contact Hours	Total Credit Hours
1	CR	CHM502	Organometallics Chemistry and Metal Clusters	4:0:0	4:0:0	4	4
2	CR	CHM504	Organic Reaction Mechanism-II	4:0:0	4:0:0	4	4
3	CR	CHM506	Quantum Chemistry and Chemical Dynamics	4:0:0	4:0:0	4	4
4	CR	CHM508	Spectroscopy:I(Techniques for Structural elucidation of Organic Compounds)	4:0:0	4:0:0	4	4
5	AEC	MAT 528/ CHM528	Mathematics for Chemists(for B.Sc. Medical students)/ Chemistry of biological systems (for B.Sc. Non Medical students)	3:0:0	3:0:0	3	3
6	SEC	PHY 540	Research Methodology &Intellectual Properties Rights	3:0:0	3:0:0	3	3
		II. Practical Subjects					
7	CR	CHM510	Organic Chemistry Practical-II	0:0:4	0:0:2	4	2
8	CR	CHM512	Physical Chemistry Practical-II	0:0:4	0:0:2	4	2
9	SEC	CHM500	Seminar	0:0:1	0:0:1	2	1
		Total				32	27

Total Credit Hours-27
Total; Contact Hours- 32

CR: Core Course
AEC: Ability Enhancement Core Course
SEC: Skill Enhancement Core Course

Semester-III

I. Theory Subjects

Sr. No.	Type of course	Sub Code	Subject Name	Contact Hours (L:T:P)	Credits (L:T:P)	Total Contact Hours	Total Credit Hours
1	CR	CHM601	Spectroscopy -II (Techniques for Structural elucidation of Physical & Inorganic Compounds)	4:0:0	4:0:0	4	4
2	CR	CHM603	Electrochemistry & Surface Chemistry	4:0:0	4:0:0	4	4
3	EC	CHM*	**ELECTIVE I** (Choose any one) G. CHM605 (Advanced coordination Chemistry) H. CHM607 (Environmental Chemistry) I. CHM609(Recent Trends in Inorganic Chemistry)	4:0:0	4:0:0	4	4
4	EC	CHM*	**ELECTIVE-II** Choose any one) G. CHM615 (Pharmaceutical Chemistry & Drug Design) H. CHM617 (Bio-Organic Chemistry) I. CHM619 (Organo-transition Metal Complexes)	4:0:0	4:0:0	4	4

II. Practical Subjects

1	SEC	CHM629	Project Part-I	0:0:8	0:0:4	8	4
2	CR	CHM611	Inorganic Chemistry Practical-III	0:0:4	0:0:2	4	2
3	CR	CHM613	Physical Chemistry Practical-III	0:0:4	0:0:2	4	2
	Total					32	24

Total Credit Hours-24
Total; Contact Hours- 32

CR: Core Course
EC: Elective Course
SEC: Skill Enhancement Core Course

Semester-IV

I. Theory Subjects

Sr. No.	Type of course	Sub Code	Subject Name	Contact Hours (L:T:P)	Credits (L:T:P)	Total Contact Hours	Total Credit Hours
1	CR	CHM602	Chemistry of Natural Products/Heterocyclic Chemistry	4:0:0	4:0:0	4	4
2	CR	CHM604	Bio-Inorganic Chemistry	4:0:0	4:0:0	4	4
3	EC	CHM*	**ELECTIVE III** (Choose any one) I. CHM606 (Instrumental Methods of Analysis) J. CHM608 (Nano-Science& NanoChemistry) K. CHM610(Green Chemistry) L. CHM618 (Photophysical Chemistry)	4:0:0	4:0:0	4	4
4	EC	CHM*	**ELECTIVE-IV** Choose any one) I. CHM612 (Industrial Chemical analysis & Quality Control) J. CHM614 (Polymer Science) K. CHM616 (Chemistry of Materials) L. CHM620 (Biofuels)	4:0:0	4:0:0	4	4
5	**AEC**	BOT001	Natural Hazards and Disaster Management	3:0:0	3:0:0	3	3

II. Practical Subjects

1	SEC	CHM630	Project Part-II	0:0:8	0:0:4	8	4
		Total				**27**	**23**

Total Credit Hours-23
Total Contact Hours- 27

CR: Core Course
EC: Elective Course
SEC: Skill Enhancement Core Course
AEC: Ability Enhancement Core Course

Summary of Scheme

Sem	L	T	P	Contact hrs/wk	Credits hrs/wk	CR	AEC	SEC	EC
I	19	0	10	29	24	20	4	2	0
II	22	0	10	32	27	20	3	4	0
III	16	0	16	32	24	12	0	4	8
IV	19	0	8	27	23	8	3	4	8
Total	76	0	42	120	98	60	10	15	16

SEMESTER I

Semester	**I**
Course Code	**CHM501**
Course Title	**Group Theory, Symmetry and Ligand Field Theory**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To introduce the concepts and importance of symmetry and group theory in solving chemical problems and transition metal complexes, orgnaometallics, inorganic chains, rings and cages.
Course Outcomes	The students will be able to: 1. Understand Coherent Knowledge of concepts and importance of symmetry and group theory to recognize and assign symmetry characteristics to molecules and objects 2. Solve chemical problems and transition metal complexes. 3. Analyze formation of inorganic chains, rings and cages and its application.

Syllabus

Unit-I

Symmetry: Symmetry elements, symmetry operations, symmetry elements commonly occurring molecules like NH3, CH4, SF6, PF5, SF4, Ni(CO)4, Fe(CO)5, matrix representation, group postulates, multiplication tables, point group determination, determination of reducible and irreducible representations, reduction formula for converting reducible representations into irreducible ones, character tables, construction of character tables for C2v, C3v (non-abelian group), use of symmetry in obtaining symmetry of orbitals in molecules, use of character table to determine which metal orbitals are used in σ and π bond formation in octahedral, tetrahedral and square planar transition metal complexes, qualitative splitting of s, p, and d orbitals in octahedral, tetrahedral and square planar fields using character tables and without the use of character tables.

Unit-II

Molecular Orbital Theory for Metal Complexes: Ligands symmetry orbitals and metal orbitals involved in molecular orbital formation in octahedral complexes, MOEL diagrams for octahedral, tetrahedral and square planar complexes showing σ and π bonding in transition metal complexes.

Chemistry of Transition Metals

Derivation of spectroscopic terms for d1 to d9 electronic configurations, correlation diagram for d2 ion in octahedral field, Splitting of d1 to d9 terms in an octahedral and tetrahedral field. Selection rules of d-d transitions. Spin-spin, orbital-orbital and spin-orbital coupling, LS and jj coupling . Comparison of CFSE values of d1 to d9 ions in terms of orbit splitting and R-S term splitting. Effect of CFSE on thermodynamic properties, lattice energy, heat of hydration heat of ligation and spinal structure.

Unit III

Free Ions in Medium and Strong Crystal Fields : Strong field configurations, transition from weak to strong crystal fields, evaluation of strong crystal field terms of d2 configuration in octahedral and tetrahedral crystal fields (using group theory), construction of the correlation energy level diagrams of d2 configuration in octahedral field, study of energy level diagrams for higher configurations, selection rules of electronic transitions in transition metal complexes, relaxation of the selection rule in centrosymmetric and non-centrosymmetric molecules, Orgel diagrams, Tanabe-Sugano diagrams. Spectra of octahedral, tetrahedral, distorted octahedral (Jahn Teller Effect) and square planner complexes spectrochemical series, charge transfer spectra.

Unit-IV

Electronic Spectra of Transition Metal Complexes: Variation of the Racah parameter, nephelauxetic effect band intensities, factors influencing band widths, Electronic spectra of octahedral and tetrahedral d1 to d9 metal ions, calculation of 10Dq and B with use of Orgel and Tanabe Sugano diagrams, low spin complexes of Mn3+, Mn2+, Fe3+, Co3+, Fe2+.

Spectra of second and third transition series, spectra of K2MoCl6 and [Rh(NH3)6] 3+ spectra of cis and trans[Co(en)2X2]+, [Mn(H2O)6]2+ , CuSO4.5H2O and its anhydrous complex, comparison of d–d band with f–f bands. Introduction to Charge Transfer Spectra**.**

Magnetic properties of transition metal ions and free ions. Effects of L-S coupling on magnetic properties. Temperature independent paramagnetism (TIP) in terms of crystal field theory CFT and molecular orbital theory (MOT). Quenching of orbital angular momentum by crystal fields in complexes in terms of term-splitting. magnetic properties of dn and fn metal ions**.** Anomalous magnetic moments, magnetic exchange coupling and spin crossover. Mixing in effect, first order and second order zeeman effects. Magnetic properties of polynuclear complexes, involving OH, NH2 and CN bridges.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Chemical Application of Group Theory	F.A. Cotton	Wiley Eastern
2	Inorganic Chemistry, 3rd edition	G. L. Miessler, D. A. Tarr	Pearson Education
3	Introduction to Ligand Field	B.N. Figgis	Wiley Eastern
4	Inorganic Electronic Spectroscopy	A.B.P. Lever	Elsevier.
5	Introduction to Magnetochemistry	A. Earnshaw	Academic Press.
6	Advanced Inorganic Chemistry	F.A. Cotton and G. Wilkinson	Wiley Inter-science

Semester	**I**
Course Code	**CHM503**
Course Title	**Organic Reaction Mechanism- I**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge of stereochemical aspects of organic compounds reactive intermediates and mechanism of general organic reactions including substitution, elimination and addition.
Course Outcomes	The students will be able to: 1. Understand Coherent Knowledge of mechanistic aspects in nucleophilic ,electrophilic substitution, addition and elimination reactions. 2. Analyze reaction conditions, products formation and mechanisms of some named reactions. 3. Apply various reaction pathways to develop new and notable organic compounds.

Syllabus

Unit-I

Stereochemistry: Elements of symmetry, chirality, projection formulae,configurational and conformational isomerism in acyclic and cyclic compounds, molecules with more than one chiral center. Threo and erythro isomers, methods of resolution, optical purity. stereogenicity, stereoselectivity, diastereoselectivity, D/L, R/S, *E*/*Z* and *cis*/*trans* configurational notations Prochirality – enantiotopic and diastereotopic atoms, groups and faces. Stereospecific and stereoselective synthesis. Asymmetric synthesis.Optical activity in absence of chiral carbon (Biphenyls, Allenes, Spiranes). Chirality due to helical shape. conformational analysis of cyclic compounds such as cyclopentane, cyclohexane, cyclohexanone derivatives, decalins, 1,2, 1,3-, 1,4-disubstituted cyclohexane derivatives and D-Glucose, effect of conformation on reactivity.

Unit-II

Nature of Bonding in Organic Reactions: Aromaticity in Benzenoid and non-benzenoid compounds. Huckel Rule, Alternant and non alternant hydrocarbons. Energy levels of Π(pi) molecular orbitals in simple systems. Annulenes, Antiaromaticity, Homoaromaticity, PMO approach.

Reaction Mechanism, Structure and Reactivity: Types of mechanisms in different reactions, thermodynamic and kinetic requirements, Kinetic and thermodynamic control in product formation. Transition states and reaction intermediates, Isotope effects, Hard and Soft Acid Base concept, Study of reactive intermediates – Types of intermediates, isolation and detection of intermediates, trapping of intermediates.

Unit-III

Aliphatic Nucleophilic Substitution: The S_N2, S_N1 and S_Ni mechanisms, mixed S_N1 & $S_N2$mechanism, SET mechanism. The neighbouring group mechanism (anchimeric assistance). Neighbouring group participation by pi and sigma bonds, Classical non-classical & phenonium

cations, Rearrangements in carbocations (general survey). Ester hydrolysis. Nucleophilic substitution at allylic, aliphatic trigonal and vinylic carbon.

Aromatic Electrophilic Substitution: The arenium ion mechanism, orientation and reactivity in mono substituted and di substituted aromatics. Energy profile diagrams. The ortho/para ratio, ipso attack, orientation in other ring systems. Gatterman-Koch reaction, Pechmann reaction, Houben – Hoesch reaction, Fries rearrangement.

Aromatic Nucleophilic Substitution: ArS_N1, ArS_N2 and ArSN via benzyne (Arynes) mechanisms. Reactivity effect of substrate structure, leaving group and nucleophile. The von Richter, Sommelet-Hauser and Smiles rearrangements.

Unit-IV

Free Radical Reactions: Type of free radical reactions, free radical substitution mechanism at an aromatic substrate, neighbouring group assistance. Reactivity for aliphatic and aromatic substrates at a bridgehead. Reactivity in the attacking radicals. The effect of solvents on reactivity.Allylic halogenation (NBS), oxidation of aldehydes to carboxylic acids, auto-oxidation. Coupling of alkynes and arylation of aromatic compounds by diazonium salts. Sandmeyer reaction. Free Radical Rearrangement. Hunsdiecker reaction, Kolbe reaction, Hydroxylation of aromatics by Fenton's reagent.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Highlights of Organic Chemistry	W.J. L. Nobel	An Advanced Text Book
2	Advanced organic chemistry part-A. 5th Ed	F. A. Carey and R. J. Sundberg	Springer (2007)
3	A guidebook to mechanism in organic chemistry, 6th Ed	Peter Sykes	Orient Longman
4	Stereochemistry conformation and Mechanism	P. S. Kalsi	New Age International
5	Stereochemistry of carbon compounds	Ernest Eliel	McGraw Hill, New York (1962).

Semester	**I**
Course Code	**CHM 505**
Course Title	**Thermodynamics: Chemical and Statistical Thermodynamics**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge of advanced classical and statistical thermodynamics. To understand behavior of activity coefficient, ionic strength, distribution Law, electro kinetic phenomena, diffusion, electric conduction, irreversible thermodynamics for biological systems
Course Outcomes	The students will be able to: 1. Understand Coherent Knowledge of different thermodynamic parameters for chemical reactions. 2. Analyze advanced classical and statistical thermodynamics. 3. Interpret irreversible thermodynamics for biological systems.

Syllabus

Unit-I

Classical Thermodynamics: Brief resume of concepts of thermodynamics, free energy, chemical potential and entropy. Partial molar properties, partial molar free energy, partial molar volume and partial molar heat content and their significances. Determination of these quantities. Concept of fugacity and determination of fugacity. Non-ideal systems: Excess functions for non-ideal solutions. Activity, activity coefficients, Debye-Huckel theory for activity coefficient of electrolytic solutions, determination of activity and activity coefficients, ionic strength. Application of phase rule to three component system, second order phase transitions.

Unit-II

Statistical Thermodynamics: Concept of distribution law, thermodynamic probability and most probable distribution, Ensemble averaging, postulates of ensemble averaging. Canonical, grand canonical and microcanonical ensembles, corresponding distribution laws (using Lagrange's method of undetermined multipliers).Partition functions: Translational, rotational, vibrational and electronic partition function,

Unit-III

Thermodynamics: First law of thermodynamics, relation between C_p and C_v; enthalpies of physical and chemical changes; temperature dependence of enthalpies. Second law of thermodynamics, entropy, Gibbs-Helmoholtz equation. Third law of thermodynamics and calculation of entropy.

Calculation of thermodynamic properties in terms of partition functions. Application of partition functions. Heat capacity behavior of solids-chemical equilibria and equilibrium constants in terms of partition functions, Fermi-Dirac statistics, distribution laws, and application to metals. Bose-Einstein statistics- distribution law and application to helium.

Unit-IV

Non Equilibrium Thermodynamics: Thermodynamic criteria for non-equilibrium states, entropy production and entropy flow, entropy balance equations for different irreversible processes (e.g., heat flow, chemical reaction etc.) transformations of generalized fluxes and forces, non-equilibrium stationery states, phenomenological equations, microscopic reversibility and Onsager's reciprocity relations, electro kinetic phenomena, diffusion, electric conduction, irreversible thermodynamics for biological systems, coupled reactions.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Thermodynamics for Chemists	S. Glasstone	East West Press, New Delhi (1950).
2	Thermodynamics for Students of Chemistry	J. Rajaram and J.C. Kuriacose	Lal Nagin Chand, New Delhi (1986).
3	Elements of Chemical Thermodynamics	L.K. Nash	Addision Wesley (1962).
4	Introduction to statistical thermodynamics	T. L. Hill	Addison-Wesley *Publishing*

Semester	**I**
Course Code	**CHM507**
Course Title	**Pericyclic reactions & Photochemistry (organic and Inorganic)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To give the knowledge of Pericyclic reactions, & photochemical reactions.
Course Outcomes	The students will be able to: 1. Understand the Basic principles of photochemical reactions, photochemistry of carbonyl compounds at different conditions. . 2. Analyze correlation diagrams method and Perturbation of molecular orbital (PMO) approach. 3. Apply Mechanistic and stereochemical aspects of thermally or photochemically driven pericyclic reactions.

Syllabus

Unit-I

Pericyclic Reactions: Molecular orbital symmetry, Frontier orbitals of ethylene, 1,3-butadiene, 1,3,5-hexatriene, allyl system, classification of pericyclic reactions FMO approach. Woodward-Hoffmann correlation diagrams method and Perturbation of molecular orbital (PMO) approach for the explanation of pericyclic reactions under thermal and photo-chemical conditions. **Electrocyclic reactions** – controtatory and disrotatory motions, 4n, 4n+2, allyl systems secondary effects.

Cycloadditions reactions – antrafacial and suprafacial additions, notation of cylcoadditions (4n) and (4n+2) systems with a greater emphasis on (2+2) and (4+2) cycloaddition-stereochemical effects and effects of substituents on the rates of cycloadditions, 1,3-dipolar cyclo-additions and cheleotropic reactions.

Unit-II

Sigmatropic Rearrangements-suprafacial and antrafacial shifts [1,2]- sigmatropic shifts involving carbon moieties retention and invertion of configuration, (3,3) and (5,5) sigma-tropic rearrangements, Claisen and Cope rearrangements, fluxional tautomerism, aza-cope rearrangements, introductions to Ene reactions, simple problems on Pericyclic reactions. Elecrocyclic rearrangement of cyclobutenes and 1,3 cyclohexadienes.

Unit-III

Photochemistry: Basic Principles of Photochemical Reactions: Photochemical laws – Franck-Condon principle, Interaction of electromagnetic radiation with matter, types of excitations, fate of excited molecule, transfer of excitation energy, radiative lifetimes, quantum yields, quenching rates and mechanisms, actinometry.

Determination of Reaction Mechanism: Classification, rate constants and life times of reactive energy states –determination of rate constants of reactions. Effect of light intensity on

the rate of photochemical reactions. Types of photochemical reactions – photodissociation, gas-phase photolysis.

Photochemistry of Alkenes: Intramolecular reactions of the olefinic bond – geometrical isomerism, cyclization reactions, rearrangement of 1,4- and 1, - dinenes.

Photochemistry of Aromatic Compounds: Isomerisations, additions and substitutions.

Unit-IV

Photochemistry of Carbonyl Compounds: Intramolecular reactions of carbonyl compounds – saturated, cyclic and acyclic, β, γ- unsaturated and α,β-unsaturated compounds, Cyclohexadienones. Intermolecular cycloaddition reactions – dimerisations and oxetane formation.

Miscellaneous Photochemical Reactions: Photo-Fries reactions of anilides. Photo-Fries rearrangement. Barton reaction. Photochemical formation of smog. Photodegradation of polymers. Photochemistry of vision.

Photochemistry of Transition Metal Complexes: Photo-reactions of complexes of Cr(III) – photo-aquation, photo-substitution and photo-racemization; Photo-substitution and photo-redox reactions of Co(III) complexes; Ru (II) polypyridyl and dinuclear Rh (I) isocyanide complexes as sensitizers; supra-molecular complexes as antenna. Applications of quenching and sensitization techniques in the identification of reactive state in coordination complexes

Photochemistry of Transition Metal Carbonyls: Photochemical substitution reactions of metal carbonyls with each other, n-donors, π-donors and ligands other than CO, Photochemical isomerization- positional isomerization, isomerization of ligands, Photochemical addition and elimination reactions – insertion into M-H, M-C and M-M bonds

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Molecular reactions and photochemistry	Chapman and Depuy	Pearson Education, Limited, 1972.
2	Synthetic Organic Photochemistry	W.H. Horsepool.	Springer
3	Molecular Reactions and Photochemistry	C. H. *DePuy* and O. L. *Chapman*	Prentice-Hall
4	Organic reaction mechanism, 3rd ed.	V. K. Ahluwalia	Narosa publishing house, New Dehli
5	Frontier orbital and Symmetry Controlled Pericyclic reactions	Ratan Kar	Books & Allied (P) Ltd
6	Pericyclic Reactions-A mechanistic Study	S. M. Mukherjee	Cambridge University Press
7	Photochemistry and Photophysics of Metal Complexes	D. M. Roundhill	Plenum Press, New York and London (1994).
8	V. Balzani and V. Carassiti	Photochemistry of Coordination Compounds	Academic Press, London (1970).
9	G. J. Ferraudi	Elements of Inorganic Photochemistry	John Wiley & Sons (1988).

10	GO. Horvath and K.L. Stevenson	Charge Transfer Photochemistry of Coordination Complexes	VCH Publishers Inc. (1993)

Semester	**I**
Course Code	**CSE551**
Course Title	**Computers for Chemists-Theory**
Type of course	Theory
L T P	3 0 0
Credits	3
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To learn the basic concepts of computer related chemistry like Huckel theory, pH titration, kinetics, radioactive decay etc.
Course Outcomes	The students will be able to: 1. Understand Coherent Knowledge of use different operating system and their tools easily 2. Apply word processing software, presentation software, spreadsheet software and latex. 3. Analyze use of computers in every field like teaching, industry and research.

Syllabus

Unit-I

Introduction To Computers And Computing: Basic structure and functioning of computers with a PC as illustrative examples. Memory, I/O devices, Secondary storage, Computer Software, Operating system with DOS as an example. Introduction to UNIX and WINDOWS. Data processing, Algorithms and flow charts.

Unit-II

Computer Programming In C/ FORTRAN: Concept of low level and high level languages, Compiler, interpreter, structure of program (header files, C pre-processor, standard library functions, etc.), Keywords, Character Set, Constants and variables, operators and Data types, Statement Input and output, Control structure such as IF or go to statement.

Unit-II

Programming in Chemistry: Development of small computer codes involving simple formulae in chemistry, such as calculation of mean, median, mode. Solution of a quadratic equation, Radioactive decay, kinetic energy, Addition, Multiplication of 3X3 matrix.

Unit-III

Use of Computer to Programmes: introduction to word processors and its features creating, editing, printing and saving documents, spell check, Adding page number, Header and Footer, Creating a table, Creating power point presentation, creating spreadsheets and use of different types of formulae, simple graphs, FOXPRO.

Unit-IV

Overview of: Information Technology (IT), Data Communication, Computer Networks (LAN, WAN and MAN), Introduction to Internet and Intranet technology and their applications, WWW, E-mail.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Computers and Common Sense	Hunt, R.; Shelley, J.	Prentice Hall.
2	Computational Chemistry	Norris, A.C.	1st edition, John Wiley & Sons, 1981.
3	Computer Programming in FORTRAN IV,	Rajaraman, V.	4th edition, Prentice Hall
4	Learn Programming in C	Dr.Kamaljeet Kaur, Anshuman Sharma	7th edition, LAKHANPAL PUBLISHERS
5	Fundamental of computer programming & information technology	Sumita Arora	3rd edition, Dhanpat Rai

Semester	**I**
Course Code	**CHM509**
Course Title	**Inorganic Chemistry Practical-1**
Type of course	Practical
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To synthesize the coordination complexes and To impart knowledge of various techniques for analysis of inorganic compounds.
Course Outcomes	The students will be able to: 1. Acquire Coherent Knowledge of analytical data forTitrimetric and gravimetric analysis of different cations and anions. 2. Understand the principles, and methodology involved in precipitations and its titrations for assaying different ions. 3. Discuss and apply the principles involved in the redox titrations and Prepare different types of inorganic compounds.

Syllabus

List of Experiments

A.Inorganic Preparations & Estimation

Preparation of Reinecke salt,Trinitrotriamine cobalt (III),
Preparation of Potassium trioxalatomangnate(III), Ferrous ammonium sulphate,
Preparation of Potassium trioxalatechromate(III).
Estimation of metal in complexes by Electronic Spectroscopy
Determine the total hardness of water.

B.Oxidation-Reduction Titrations

Standardization with sodium oxalate of $KMnO_4$ and determination of Ca^{2+} ion.
Standardization of ceric sulphate with Mohr's salt and determination of Cu^{2+}, NO_3 -1 and C_2O_4 ions.
Standardization of $K_2Cr_2O_7$ with Fe^{2+} and determination of Fe^{3+} (Ferric alum)
Standardization of hypo solution with potassium iodate / $K_2Cr_2O_7$ and determination of available Cl_2 in bleaching powder, Sb^{3+} and Cu^{2+}.
Determination of hydrazine with KIO_3 titration.

C.Precipitation Titrations

$AgNO_3$ standardization by Mohr's method by using adsorption indicator.
Volhard's method for Cl^- determination.
Determination of ammonium / potassium thiocyanate.

D.Complexometric Titrations

Analysis of two cation-system using EDTA
Determination of Cu^{2+} and Ni^{2+} by using masking reagent by EDTA titration.
Determination of Ni^{2+} (back titration).
Determination of Ca^{2+} (by substitution method).

E. Gravimetric Analysis

Determination of Ba^{2+} as its chromate.
Estimation of lead as its lead molybdate.
Estimation of chromium (III) as its lead chromate.
Estimation of Cu^{2+} using Ammonium/ Sodium thiocyanate.
Note: Perform at least any two from each section.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Vogel's quantitative analysis 6 Edn	Mendham, Denny	Pearson Education 2002
2	Synthesis and Technique in Inorganic chemistry	G. S.Girlomi; R.J. Angleci	3rd edn.; University Science Books.
3	Advanced Practical Inorganic Chemistry	Ayodha Singh	Campus Books 2002

Semester	**I**
Course Code	**CHM511**
Course Title	**Organic Chemistry Practical-1**
Type of course	Practical
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To learn the basic organic preparations and organic reagents like reducing agents and oxidizing agents.
Course Outcomes	The students will be able to: 1. Adopt safe laboratory practices by handling laboratory glassware, equipment, and chemicals. 2. Understand the basic nature of reagents like reducing agents and oxidizing agents. 3. Apply & propose starting materials, functional groups, mechanism, and typical reaction conditions.

Syllabus

List of Experiments

Synthesis: Synthesis, purification and identification of organic compounds by recrystallization/functional group identification:

i. Oxidation: Adipic acid from cyclohexanol
ii. Aldol condensation: Dibenzal acetone from benzaldehyde
iii. Sandmeyer reaction: p-Chlorotoluene from p-toluidine
iv. Cannizzaro reaction: Benzyl alcohol and benzoic acid from benzaldehyde
v. Aromatic electrophilic substitutions: p-nitroaniline from aniline
vi. Aromatic electrophilic substitutions: Picric acid from phenol
vii. Beckmann Rearrangement: Benzanilide ← Benzophenone oxime ← Benzaldehyde
viii. Reduction: Benzhydrol from benzophenone [$NaBH_4$ reduction]
ix. Esterification: Methyl benzoate from benzoic acid
x. Carbohydrate Modification: Osazone derivative from carbohydrates
xi. Haloform reaction: Iodoform synthesis from acetone / ethyl alcohol
xii. Sublimation: Synthesis/purification of Phthalic anhydride from Phthalic acid
xiii. Preparation of p- Iodonitrobenzene from p-nitroaniline.
xiv. Preparation of benzyl alcohol and benzoic acid (Cannizzaro's reaction).
xv. Preparation of Dibenzal acetone from benzaldehyde (Claisen-Schmidt reaction).
xvi. Preparation of Acetanilide, bromoacetanilide, bromoaniline.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Experimental Organic Chemistry,	Harwood, L.M., Moody, C.J.	1st edition, Blackwell Scientific Publishers, 1989.
2	Text Book of Practical Organic Chemistry	Vogel, A.I.	ELBS, IVth edition, Longman Group Ltd., 1978.

3	Practical Organic Chemistry	. Mann, F.G.; Saunders, B.C.	4th edition, New Impression, Orient Longman Pvt. Ltd., 1975

Course Code	**CSE553**
Course Title	**Computer in Chemistry Lab-I**
Type of Course	Lab
L T P	0 0 1
Credits	1
Course Prerequisites	Knowledge of C ,C Programming Language
Course Objectives (CO)	Allows the students to know about background functioning of System Programs
Course Outcomes	The students will be able to: 1. Understand about background functioning of System Programs. 2. Use working of the internet for the use of domains, IP addresses, URLs and different web browsers. 3. Acquire knowledge to search information using search engines for different programme.

SYLLABUS

1.Word Processor software :Word: To familiarize with parts of Word window, To create and save a document ,To set page settings, create headers and footers, To edit a document and resave it To use copy, cut and paste features. To create a table with specified rows and columns To create a table with specified rows and columns, To select a table, a row, a column or a cell ,To inset new row and/or a column, To delete a row and/or a column.

*Excel:*To familiarize with parts of Excel window, To create and save a workbook with single and/or multiple worksheets To edit and format text as well numbers To insert new row and/or column in a worksheet,'To delete a row and/or column in a worksheet.

Power point: To familiarize with parts of PowerPoint window create and save a new presentation, To apply design templates to a presentation insert, edit and delete a slide, To use different views of slides . To use slide show from beginning or from the current slide To preview and print a presentation. To check spellings in a presentation, To add clip art and pictures in a slide, To add chart, diagram and table in a slide, To set animation for a selected slide and/or for entire presentation.

2 Exploring the Internet: To understand the working of the internet that include the use of protocols, domains, IP addresses, URLs, web browsers, web servers, mail-servers, etc. create email-account, sending , mails, receiving mails, sending files as, attachments, etc. To login to a remote computer, To search information using search engines.

Write a Program to display a message.
Write a program to display greater of two numbers.
Write a program to find area of a circle.
Write a program to find addition of two numbers.
Write a program to convert Celsius temperature to Fahrenheit.
Write a program to implement calculator in c
Write a program to find factorial of a number.

Write a program to implement and print an array elements in C.
Write a program in C to print two dimensional array.

Text and Reference Books:

Sr. no.	Name	Author(S)	Publisher
1	ANSI C	Balagurusamy	McGraw Hill Education India Pvt Ltd
2	Let us C	Yeshavant Kenetkar	BPB Publication

SEMESTER

II

Semester	**II**
Course Code	**CHM502**
Course Title	**Organometallics Chemistry and Metal Clusters**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject.
Course Objective	The main goal of this branch is to study the organometallics, π-acid ligands, Inorganic Rings, Chains and Metal Cluster which containing at least one bond between a metal and carbon
Course Outcomes	The students will be able to: 1. Explain basic properties,formation, reaction mechanism of organometallic compound. 2. Understand synthesis, properties, bonding and structures of organometallic compound. 3. Understand the principles behind the formation of metal cluster compounds , stability and application of Inorganic Rings, Chains and cages.

Syllabus

Unit -I

Organic-transition metal chemistry: The basis of 18_{e^-} Rule, 16-electron rules , Exceptions to eighteen electron rule, isolobal analogy.

complexes with π-acceptor and σ-donor ligands, π-acid ligands: π-acceptor character of CO, O_2, N_2, NO, PH_3 molecules in terms of MOEL diagram,

Organometallics: Methods of preparation in perspective-organo lithium compounds: structure and bonding & reaction of carbolithiatic organometallics of group 2 and 12 e.g. Mg and Zn, Cd and Hg: Preparation and structure of organoaluminium compounds, Technical applications of tris (alkyl) aluminium compounds.

Molecular orbital description of ligands to transition metals. Description of bonding of two electron ligands to Transition Metals. Preparation of olefin Transition Metal Complexes, olefenic and acetylenic complexes, 2π – ligands: Allylic and 4π – complexes of cyclopentadiene. Synthesis and structure. 4π–ligands: Butadiene, cyclobutadiene, 2π–complexes of cyclopentadiene, pentadiene, cyclic dienes, cyclopentadienyls, Molecular orbital picture of bonding in ferrocene.

Unit-II

Homogeneous hydrogenation of unsaturated compounds, dihydrocatalysis, monohydrido compounds, selected application of dihydride oliphinic hydrogenation catalyst, Wilkinsons catalyst.

Hydrosilation of unsaturated compounds, aldehydes and ketones, Ziegler-Natta polymerization of ethylene and propylene, water gas shift reaction, acetic acid synthesis by carbonyls

Unit-III

Reaction by electrophilic attack on Coordinated ligands: electrophilic cleavage of metal carbon σ bond, mechanism for alkyl complexes of d electron, electrophilic modification of coordinated ligands: attack at α position, attack at β position, attack at γ position.
Synthetic application of transition metal carbonyl compound: Synthesis and important reaction of metal carbonyls, coupling reaction, carbonylation by neutral metal carbonyls, carbonylation by anionic metal complexes, carbonylation by metal acyl complexes, decarbonylation of aldehydes and acid chlorides.

Unit-IV
Inorganic Rings, Cages and Metal Cluster: Inorganic cages

Rings: Synthesis and application of Borazines, Phosphazenes and other heterocyclic inorganic ring, homocyclic inorganic systems.
Cage : cages of P and S, oxides & sulphides, metal carbonyl and halide clusters, compounds with metal-metal multiple bonds, boron cage compounds, boranes, carboranes and metallocene carboranes,– synthesis and structure, .
Metal cluster: metal-metal bond, binuclear compounds, trinuclear clusture, fluxional molecules.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Organometallics: A Concise Introduction	C. Elschenbroich and A. Salzer,	2nd Ed., VCH1992.
2	Inorganic Chemistry Principles of Structure and Reactivity	J.E. Huheey	Harper Interscience
3	Wilkinson, Advanced Inorganic Chemistry	F.A. Cotton and G. Wilkinson	Ed. V & VI.Wiley Inter-science.
4	Inorganic Chemistry	G. L. Miessler, D. A. Tarr	3rd edition, Pearson Education

Semester	**II**
Course Code	**CHM504**
Course Title	**Organic Reaction Mechanism-II**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To impart knowledge of stereochemical aspects of mechanisms of organic reactions viz E2, E1, E1cb, addition to carbon-carbon, carbon-hetero multiple bond and to understand different rearrangements and oxidations and reductions.
Course Outcomes	The students will be able to: 1. Propose and determine the mechanism and feasibility of a chemical reaction. 2. Apply mechanistic aspects in nucleophilic and electrophilic substitution. 3. Interpret reaction conditions, products formation and mechanisms of some addition and rearrengement reaction.

Syllabus

Unit-I

Elimination Reactions: The E2, E1, E1cB mechanisms. Orientation of the double bond. Effects of substrate structure, attacking base, leaving group and medium on reactivity. Mechanism and orientation in pyrolytic eliminations.

Addition to Carbon – Carbon Multiple Bonds: Mechanistic and stereochemical aspects of addition reactions involving electrophiles, nucleophiles and free radicals, regio and chemoselectivity, orientation and reactivity. Addition to cyclopropane ring.

Unit-II

Addition to Carbon – Hetero Multiple Bonds: Mechanism of metal hydride reduction of saturated and unsaturated carbonyl compounds, acids, esters and nitriles. Wittig reaction. Mechanism of condensation reactions involving, Kneoevenagel, Mannich, Benzoin, Perkin and , Reformatski reaction.

Formation of Carbon-carbon Bond: Principle, disconnections and synthons, electrophilic and nucleophilic carbon species. Base catalyzed condensations; Aldol condensation, Claisen reaction, Perkin reaction, Stobbe condensation, Darzen condensation, Use of malonic, acetoacetic and cyanoacetic esters, Micheal addition, Wittig reactions. Use of acetylides, Acid-catalyzed condensation – self condensation of olefins, Friedal-Craft's reactions, Fries reactions

Unit-III

Rearrangements: General mechanistic considerations – nature of migration, migratory aptitude, memory effects. A detailed study of the following rearrangements: Pinacol-pinacolone, Wagner-Merwein, Demjanov, Benzil-Benzilic acid, Favorskii, Arndt-Eistert synthesis, Beckmann, Hofmann, Curtius, Schmidt, Baeyer-Villiger, Shapiro reaction, Cope, Gabriel–Colman, Smiles and Sommelet–Hauser rearrangements

Unit-IV

Oxidation-I: Introduction**.** Different oxidative processes. Hydrocarbons **-** alkenes, aromatic rings, saturated CH groups (activated and unactivated). Alcohols, diols, aldehydes, ketones and carboxylic acids. Amines, hydrazines, and sulphides.

Oxidation Reactions-II : Oxidations with ruthenium tetraoxide, iodobenzene diacetate and thallium (III) nitrate,DDQ, PCC, CAN, selenium dioxide, peroxyacids, DCC. Oxidation reactions with special emphasis on Baeyer-villeger reaction, Cannizarro oxidation-reduction reaction

Reduction: Introduction**.** Different reductive processes. Hydrocarbons **-** alkanes, alkenes, alkynes and aromatic rings. Carbonyl compounds – aldehydes, ketones, acids and their derivatives. Epoxides. Nitro, nitroso, azo and oxime groups. Hydrogenolysis. Sodium borohydride, sodium cyano borohydride, LAH, disobutyl aluminium hydride, tin hydride, trialkyl tin hydride, trialkyl silanes, alkoxy substituted LAH, DIBAL, diborane, diisoamyl borane, hexyl borane, 9-BBN, isopinocamphenyl and disiopinocamphenyl borane. Reduction reactions with particular emphasis on Wolf-Kishner reduction, Clemensen reduction.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Principles of Organic Synthesis	Norman and Coxon	*CRC Press*
2	Advanced Organic Chemistry	Jerry March	Wiley-Interscience; 3rd edition (March 11, 1985)
3	Advanced Organic Chemistry	F.A. Carey, R.J. Sunberg	Springer; 5th edition (May 27, 2008)
4	Highlights of Organic Chemistry	W, J.L. Nobel	Advanced Text Book.

Semester	**II**
Course Code	**CHM506**
Course Title	**Quantum Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To acquire knowledge of the Quantum theory & quantum chemical description of chemical bonding, reactivity and their applications in molecular spectroscopy and inorganic chemistry.
Course Outcomes	The students will be able to: 1. Acquire Knowledge about Electronic and Hamiltonian operators for molecules. 2. Understand Quantum chemical description of angular momentum. 3. Use Quantum chemical description of chemical bonding, reactivity and their applications in molecular spectroscopy and inorganic chemistry.

Syllabus

Unit-I

Quantum Theory: Introduction and Principles: Black body radiations, Planck's radiation law, photoelectric effect, Compton Effect, De- Broglie hypothesis, the Heisenberg's uncertainty principle, Rydberg relation for explaining atomic spectrum of hydrogen. Bohr's Theory and its limitation solution of classical wave equation by separation of variables method, operators and observations, normal and orthogonal functions, Hermitian and unitary operators, introduction to differentiation and integration, Eigen value equation. Hamiltonian operator, interpretation of wave function, postulates of quantum mechanics.

Unit-II

Applications of Quantum Postulates: Solution of particle in one and three dimensional box, degeneracy, the linear harmonic oscillator, rigid rotators, quantization of vibrational and rotational energy levels, hydrogen and hydrogen like atoms.

Angular Momentum: Commutative laws, need of polar coordinates, transformation of Cartesian coordinate into polar coordinate, angular momentum of one particle system, orbital angular momentum, the ladder operator method for angular momentum, spin angular momentum and their relations

Unit-III

General Orbital Theory of Conjugated Systems: Chemical bonding, linear combination of atomic orbital, overlap integral, coulomb's integral, bond order, charge density calculations for ethylene, allyl system, butadiene system, cyclo butadiene cyclo propenyl system.

Unit-IV

The Approximate Methods: Need for approximation methods, Perturbation and Variation methods and their application to Helium atom.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Physical Chemistry, A Molecular Approach	MacQuarrie and Simon	University Science Books, 1997
2	Quantum Chemistry	Ira N. Levine	Prentice Hall
3	Physical Chemistry	P. W. Atkins	ELBS, Oxford, 1997.
4	Quantum Chemistry	H. Eyring, Kimball and Walter	Nabu Press (September 12, 2011)

Course Code	**CHM508**
Course Title	**Spectroscopy:1(Techniques for Structural elucidation of Organic Compounds)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To get familiarized with various spectroscopic techniques such as UV, IR, NMR and Mass spectroscopy and illustrate their application for structural elucidation of organic molecules.
Course Outcomes	The students will be able to: 1. Acquire Coherent and advanced knowledge of the principles and techniques in spectroscopy. 2. Understand electronic,Vibrational, proton NMR & 13C NMR and mass spectrometry methods of analysis. 3. Apply spectroscopic methods (UV,IR,1H-NMR,13C-NMR& mass spectrometry) in organic structure elucidation.

Syllabus

UNIT-I

Introduction to spectroscopy, Nature of radiation. Energies corresponding to various kinds of radiation, Experimental techniques, intensities of spectral lines, Selection rules and transition moments, Linewidths, Broadening.

Ultraviolet and Visible Spectroscopy: Various electronic transitions (185-800 nm), Beer-Lambert law, Factors affecting electronic transitions: effect of solvent, Steric effect in biphenyls.**U**ltraviolet bands for carbonyl compounds, unsaturated carbonyl compounds, dienes, conjugated polyenes. Woodward- Fieser-rules for conjugated dienes and carbonyl compounds & Benzanoid systems, Ultraviolet spectra of aromatic and heterocyclic compounds.

UNIT-II

Infrared Spectroscopy: Introduction, Dispersion IR spectrometer and FT IR spectrometer: Instrumentation and sample handling. Characteristic vibrational frequencies of alkanes, alkenes, alkynes, aromatic compounds, mononuclear and polynuclear aromatic compounds, alcohols, ethers, phenols and amines. Detailed study of vibrational frequencies of carbonyl compounds (ketones, aldehydes, esters, amides, acids, anhydrides, lactones, lactams and, conjugated

carbonyl compounds). Effect of hydrogen bonding and solvent effect on vibrational frequencies, overtones, combination bands and Fermi resonance. FT IR. IR of gaseous, solids and polymeric materials.

Nuclear Magnetic Resonance Spectroscopy I: The nuclear spin, precessional motion. Larmor frequency, Energy transitions, spin – spin and spin –lattice relaxation, measurement techniques (CW and FT methods). Solvent used, reference standards. Chemical Shift, Simple applications of chemical shift, ring currents and aromaticity, anisotropy shifts of 1H and 13C,inductive effect, intermolecular forces affecting the chemical shifts. shielding constant, isotopic nuclei, integration of signals 1H- NMR spectra: spin-spin coupling, coupling constants, Karplus relationship., interpretation of spectra.

UNIT-III

Nuclear Magnetic Resonance Spectroscopy II: chemical shift values and correlation for protons bonded to carbon (aliphatic, olefinic, aldehydic and aromatic) and other nuclei (alcohols, phenols, enols, carboxylic acids, amines, amides & mercapto), chemical exchange, First and second order spectra A2, AB, AX, AB2, AX2,A2B2, A2X2, ABX and ABC spin systems with their coupling constants, Interaction between two or more nuclei, splitting due to vicinal and Germinal protons. Long range coupling, Decoupling Techniques, double resonance and shift reagent methods, Effect of deuteration, Effects of chemical exchange, fluxional molecules, Hindered rotation on NMR spectrum.

13C NMR: Introduction, nuclear overhauser enhancement (NOE), 2D NMR, Correlation spectroscopy (COSY), Homo COSY (1H-1H COSY), Hetro COSY (1H-13C COSY, HMQC), long range 1H-13C COSY (HMBC), NOESY, DEPT techniques, 13C NMR spectra, their interpretation and applications.

UNIT-IV

Mass Spectrometry: Introduction, ion production - EI, CI, FD and FAB, factors affecting fragmentation, ion analysis, ion abundance, molecular ion peak, metastable peak, McLafferty rearrangement. Nitrogen rule. Mass spectral fragmentation of organic compounds, common functional groups, High resolution mass spectrometery. Examples of mass spectral fragmentation of organic compounds with respect to their structure determination

The combined application of UV, IR, NMR and mass spectra in solving advanced spectroscopic problems.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Instrumental Methods of Chemical Analysis	G.W.Ewing,	McGraw Hill Pub, 1975.
2	Spectrometric Identification of Organic Compounds	Robert M. Silverstein	7th edition John Wiley
3	Organic Spectroscopy.	W. Kemp	Macmillan; 2nd edition (1987);
4	Organic Spectroscopy: Principles and Applications	Jag Mohan	Himalaya ***Publishing*** House, Bombay, 1 992.
5	Applications of absorption spectroscopy of organic compounds. (1965).	Dyer, J. R.	Phi Learning.

6	Spectroscopy of organic compounds. (2007).	Kalsi, P. S.	New Age International.
7	Introduction to solid NMR Spectroscopy. (2010)	Melinda, J.D.	Wiley India Pvt Ltd
8	Introduction to spectroscopy. (2008).	Pavia, D. L., Lampman, G. M., Kriz, G. S., and Vyvyan, J. A.	Cengage Learning.

Course Code	**MAT 528**
Course Title	**Mathematics for Chemists(for B.Sc. Medical students)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To learn basic concept of chemistry numerical like order of reaction, method of partial fractions, area under a curve etc.
Course Outcomes	The students will be able to: 1. Understand basic numerical methods like order of reaction, method of partial fractions 2. Analyze & interpret the area under a curve using Integral Calculus. 3. Apply solution of linear equations by using Determinants and Matrices.

Unit-I

Trigonometry: Definition of sin, cos, tan, cot, sec, cosec functions with the help of unit circle, values of sin x cos x cos x for x = 0, n/6, n/3, n/2. Meaning of a trignometrical identity. The following identities (no need of derivation and proof. However, application has to be emphasized).

Cos2 x+sin2 x = 1,Sin (x=2)-Sin2 x/½ sin x, Cos (x-2)Cos x, Cos(-x = cos x; sin(-x) = sin x Sin(-x) = sin x ; cos (-x) = --cos x, Sin(+x)= ---xin x; (+x) = --co x, Sin 2x = 2 sin x cos x Cos 2x = 2 sin x cos x, Cos 2x = 2lcosx-1, Tan (x) = -- sin x; cos x,Tan (x) = -- tan x; tan(x/2 – x = cot x), Tan (-x) = -tan x, Tan 2x = 2 tan x/(1-tan2 x)

Unit-II

Determinants and Matrices: Definition and expansion properties of determinants, product of two determinants of 3rd order. Introduction to various terms Matrix, row, column, diagonal unit. Sub, square, equal matrices, null, symmetricular, order of, character of, transpose of, adjoint of, inverse of matrices. Addition multiplication, Multiplication of matrices.

Unit -III

Differential Calculus : Differentiation of standard functions, theorems relating to the derivative of the sum, difference, product and quotient of functions, derivative of trigonometric functions, inverse trigonometric functions, logarithmic functions and exponential functions, differentiation of implicit functions, logarithmic differentiation.

Unit-IV

Integral Calculus : Integration as an inverse of differentiation summation, area under a curve, indefinite integrals of standard forms, method of substitution, method of partial fractions, integration by parts, definite integrals.

Text and Reference Books:

S. No	Name	Author(S)	Publisher
1	Differential Calculus	Santi Narayan	S. Chand And Company
2	Integral Calculus.	Santi Narayan	S. Chand And Company
3	Higher Engineering Mathematics	B.S. Grewal	Khanna Publishers

4	Mathematical Techniques in Chemistry	Joseph B. Dence	Wiley; 1St Edition edition (1975)
5	Mathematics of Physics and Chemistry	Margenau and Murphy	Van Nostrand
6	A Text Book of Engineering Mathematics	B.L. Moncha and H.R. Choudhary	Young Press; 2 edition

Semester	**II**
Course Code	**CHM 528**
Course Title	**CHEMISTRY OF BIOLOGICAL SYSTEMS (For Non-Medical Students)**
Type of course	Theory
L T P	3 0 0
Credits	3
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To impart knowledge of molecular structure and interactions present in various bio-molecules that assist in functioning and organization of biological cell.
Course Outcomes	The students will be able to: 1. Acquire basic knowledge about organization and working principles of various components of living cell. 2. Understand basic principles of structure, function, and folding of biomolecules 3. Acquire knowledge of molecular structure and interactions of proteins, carbohydrates, lipids and nucleic acids.

Unit 1

Introduction: Cell structure and functions, general characteristics of cells, difference between prokaryotic and eukaryotic cells, difference between plant and animal cells, cell organelles. Tissues, organs and organ systems.

Thermodynamics and kinetics of biological processes, Water – physical properties and structure of water molecules, Role of water in life, pH, Acidic and basic buffers, Biological buffers, solution equilibria, Henderson-Hasselbalch equation, Hofmeister series, Chaotropic and kosmotropic ions/co-solvents.

Unit 2

Amino Acids and Peptides: Classification and properties of amino acids, Isoelectric point, Separation of amino acids, peptide and polypeptides, primary structures, structure of peptide bond, synthesis of peptides, N-terminal, C-terminal and sequence determination.

Carbohydrates: Biologically important monosaccharides, disaccharides and polysaccharides, glycoproteins, role of sugars in biological recognition. An overview of Metabolism: citric acid cycle, glycolysis.

Unit 3

Proteins: Secondary structure of proteins with emphasize on supramoelcular --sheets, supersecondary structure and triple helix structure of collagen, tertiary structure of protein-folding, quaternary structure of protein, Automated Peptide synthesis, in-vivo and in-vitro protein folding, protein misfolding and conformational diseases.

Unit 4

Nucleic Acids: Purine and pyrimidine bases, Nucleotides, Nucleosides, base pairing via H-bonding, DNA, RNA structure and conformation, double helix model of DNA, different types of RNA and their functions, the chemical basis for heredity, overview of replication of DNA, transcription, translation and genetic code, ATP.

Lipids: Lipid classification, lipid bilayers, liproproteins-composition. high density (HDL) and low-density (LDL) lipoproteins and function, Phospholipids, membrane proteins - integral membrane proteins, .

Text and Reference Books:

S. No.	Name/Title	Author	Publisher
1	Cord Biology		South Western Educational Publications, Texas, 2000.
2	Contemporary Enzyme kinetics and Mechanisms	D. L. Purich	Academic Press, 1983.
3	Bio-organic Chemistry, A chemical approach to enzyme action	Dugas H	Springer 2003.
4	Text book of Organic Medical and Pharmaceutical Chemistry,10th Ed	Wilson, Gisvold & Dorque	Lippincoh pover publishers, 1998.
5	Organic Chemistry, Third Edition,	Paula Yurkanis Bruice	Pearson Education

Course Code	**PHY540**
Course Title	**Research Methodology and Intellectual Property Rights**
Type of course	Theory course
L T P	3 0 0
Credits	3
Course prerequisite	B. Sc. Medical or Non-medical
Course Objective	Student will be understand to how to identify a research problem, know the importance of educational research and role of Intellectual Property Rights (IPR) in research and development
Course Outcomes	The students will be able to: 1. Acquire & Understand significace of IPR, copyright laws in present scinario. 2. Identify a research problem, educational research, interpretation of the results and report writing. 3. Apply role of Intellectual Property Rights (IPR) in research and development.

UNITI

Research Methodology: Types and method of research, Research process; criteria of good research. Defining and formulating the research problem, selecting the problem, necessity of defining the problem, importance of literature review in defining a problem, Meaning of research

design; need for research design; important concepts related to research design; different research designs. Writing research proposal: Characteristics of a proposal; content and organization of a proposal.

UNIT II

Interpretation and report writing: Meaning of interpretation; technique of interpretation; precautions in interpretation; significance of report writing; layout of research report; types of reports; Organization and writing of research paper, Presentation of research work-oral, poster and writing of research paper; Precautions for writing research report, Application and uses of common softwares in chemistry and physics.

UNIT III

IPR: Introduction and the need for intellectual property right (IPR) - Kinds of Intellectual Property Rights: Copyright, protection under copyright law, rights, transfer of copyright, infringement,

Trademarks its objectives, types, rights, protection of goodwill, infringement, passing off, Defenses, Domain name, trade secrets. Design, Geographical Indication.

Introduction to the leading International Instruments concerning Intellectual Property Rights: the Berne Convention, Universal Copyright Convention, The Paris Convention, Patent Co-operation Treaty, TRIPS, The World Intellectual Property Organization (WIPO) and the UNESCO. Infringement. IPR in Pharmaceuticals and drug designing

UNIT IV

Ethical issues: Citation and acknowledgement, Reproducibility, Review of published research in the relevant field, plagiarism.

Patent and Patents Writing, Patent Act 1970 and its amendments. Procedure of obtaining patents, Chemical safety and ethical handling of chemicals. Safety rules of laboratory acquaintance of experimental set up, importance of safety and security of data.

Industrial Designs its objectives, rights, registration, infringements, and Defenses of Design, Need for Protection of Industrial Designs, The Designs Act, 2000

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Research Methodology: Methods & Techniques (Rev. Ed.)	C.R. Kothari	New Age International. New Delhi
2	An Introduction to Research Methodology	B.L. Garg, R. Karadia, R., F. Agarwal, F. and U.K. Agarwal	RBSA Publishers
3	Qualitative Inquiry and Research Design: Choosing Among Five Approaches	John W. Creswell	SAGE Publication
4	Principles of Intellectual Property	N.S. Gopalakrishnan, and T.G. Agitha	Eastern Book Company
5	Law relating to patents, trade marks, copyright designs and geographical indications	B.L.Wadehra	Universal Law Publishing
6	An Introduction to Intellectual Property Rights	Venkataraman M	Affliated East-West Press

Semester	**II**
Course Code	**CHM510**
Course Title	**Organic Chemistry Practical- II**
Type of course	Laboratory Course
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	Organic synthesis and quantitative analysis of organic compounds.
Course Outcomes	The students will be able to: 1. Acquire basic knowledge of organic synthesis of organic compounds. 2. Analyze & Interpret the quantitative analysis of organic compounds. 3. Propose methodologies for the extraction of Organic Compounds from Natural Sources.

1. Preparation, separation and purification of organic compounds, and their characterization by spectral techniques (UV, IR, PMR, CMR and MS)

1. Synthesis of 2-chloro-4-bromoaniline from aniline (Bromination and chlorination)
2. Synthesis of methyl orange from aniline. (Aromatic electrophilic substitution and diazocoupling).
3. Synthesis of benzpinacol and its pinacol rearrangement.
4. Synthesis of o-chlorobenzoic acid from phthalimide. Synthesis of acridone from o-chlorobenzoic acid. (Hofmann bromamide and Sandmeyer's reaction).
5. Synthesis of 2, 4-dinitrophenyl hydrazine from chloro benzene. (Electrophilic and nucleophilic substitution reactions on aromatic ring).
6. Synthesis of triphenylcarbinol from bromobenzene. (Grignard reaction)

2. Quantitative Analysis of Organic Compounds:

1. Estimation of phenol/aniline using bromate-bromide solution. (The application to find the purity of the sample and to determine the amount in given solution).
2. Determine the number of hydroxyl and amino groups in the given sample by the acetylation method.
3. Determine the mol. wt. of the given ketone by using 2,4-DNP method.
4. Estimation of reducing sugar by Fehling solution method.
5. To determine the sponification value of the given fat or oil sample.
6. To determine the iodine number of the given fat or oil sample.

3. Chromatography:

Separation and identification of the sugars present in the given mixture of glucose, fructose and sucrose by paper chromatography and determination of R_f values

4. Synthesis using Ionic liquids

Preparation of ionic liquid.

Synthesis of chiral compounds using ionic liquids and determination of optical activity of the product by polarimeter

5. Extraction of Organic Compounds from Natural Sources

(i) Isolation of caffeine from tea leaves.

(ii) Isolation of casein/lactose from milk

(iii)Isolation of piperine from black pepper

(iv)Isolation and purification of lecithin from Soyabean seeds

6. Synthesis using Microwave/Sonicator

(i) Alkylation of diethyl malonate with benzyl chloride.

(ii) Synthesis of Heterocyclic compounds using multi component reactions

Note: Perform at least any two from each section.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Laboratory Experiments in Organic Chemistry	R. Adams, J. R. Johnson and C. F. Wilcox.	The Macmillan Limited, London.
2	An Introduction to Modern Experimental Organic Chemistry	R. M. Roberts, J. C. Gilbert, L.B.Rodewald and A. S. Wingrove Holt	Ranehart and Winston Inc. New York.
3	Introduction to Organic Laboratory Techniques – A Contemporary Approach	D. L. Pavia, G. M. Lampmana and G. S. Kriz	W. B. Saunders Company, 1976.

Semester	**II**
Course Code	**CHM512**
Course Title	**Physical Chemistry Practical- II**
Type of course	Laboratory Course
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To impart knowledge and hand-on experiences of different analytical and thermodynamic techniques (conductometry, pHmetry and other viscosity and surface tension measurements) for chemical and biomolecular analysis
Course Outcomes	The students will be able to: 1. Acquire basic knowledge about analytical techniques such as conductometric, pH metric and potentiometry techniques. 2. Understand and Apply different thermodynamic techniques like viscosity and surface tension measurements for solutions. 3. Analyze determination of solubility of different inorganic and organic salt.

1. Viscosity:

(i) To determine the coefficient of viscosity of given liquid by Ostwald's viscometer.

(ii) Determination of relative and absolute viscosity of a given liquid.

(iii) Determination of percentage composition of a liquid mixture by viscosity measurement.

(iv) Determination of molecular weight of a high polymer (say polystyrene) by viscosity measurement.

2. Surface Tension:

(i) Determination of surface tension of given liquid by drop no. method by stalgmometer.

(ii) To determine the C.M.C. of a soap (sodium or potassium lauryl sulphate by surface tension measurements and to compare cleansing power of two detergents.

(iii) Determination of surface tension of alcohols & Determination of Parachor value of $>CH_2$ group.

(iv) To measure interfacial tension and to test the validity of Antonoff's rule.

3. **Solubility:**

(i)Determination of solubility of inorganic salt in water at different temperatures and hence to draw the solubility curve.

(ii)Determination of heat of solution of a substance by solubility method

(iii) To study the effect of addition of an electrolyte on the solubility of an organic acid.

(iv) To study the variation of solubility of Ca $(OH)_2$ in NaOH solution and hence determine the solubility product.

(v)Determine the solubility (g/litre) of sparingly soluble lead sulphate from conductance measurements.

(vi)To obtain the mutual solubility curve of phenol + H2O, and hence the Upper Consolute point,

4. Colloidal State:

(i) To compare the precipitation power of Na^+, Ba^{+2} & $A1^{+3}$ ions for As_2S_3 sol.

(ii) To study interaction between arsenious sulphide and ferric hydroxide sol.

5. Potentiometric/ conductometric titrations:

(i) Preparation of buffers and measurement of their pH.

(ii) To determine the strength, dissociation constant of given acid pH metrically.

(iii)Titration of weak acid /Weak base conductometrically.

(iv)Titration of strong acid /strong base conductometrically.

(v)To determine dissociation constant of given acid conductometrically.

(vi)Compare the relative strength of CH_3COOH and $ClCH_2COOH$ from conductance measurements.

(vii)Titrate a given mixture of HCl and CH_3COOH against NaOH solution conductometrically

(viii)Determine the dissociation constant of acetic acid in DMSO, DMF and dioxane by titrating it with KOH.

(ix)Determine the activity coefficient of an electrolyte at different molalities by e.m.f. measurements.

Note: Perform at least any three from each section.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Experimental Physical Chemistry	Arthur M. Halpern, George C. McBane	Freeman,2006.
2	Experiments in Physical Chemistry, 5th ed.,	Schoemaker et al.	MGH, 1989
3	Chemistry Experiments for Instrumental Methods	Sawyer, Heineman, Beebe	Wiley, 1984 .
4	Physical Chemistry Practical.	Maity S., and Ghosh, N.(	New Central Book Agency (P) Ltd. 2012).
5	Senior Practical Physical Chemistry.	Khosla, B.D., Garg, V.C., and Gulati A.R.	S. Chand and Sons. (2007).
6	Advanced Practical Physical Chemistry.	Yadav, J. B.	Krishna Prakashan Media. (2006).
7	Experiments in Physical Chemistry,	Ghosh, J.C.	Bharati Bhavan. (1990).

Semester	**II**
Course Code	**CHM500**
Course Title	**Seminar**
Type of course	Laboratory Course
L T P	0 0 2
Credits	1
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	The course would develop soft skills of students, scientific aptitude, critical thinking, research writing and research presentation.
Course Outcomes	The students will be able to: 1. Investigate various aspects related to the chemistry. 2. Appreciate the literature and its relevance to his/her topic of interest how to write a report on a given topic. 3. Technical write and presentation on a given topic of research and commercial worth of chemistry.

The seminar must include discussion on topics such as awareness about weapons of mass destruction (chemical, biological, radiological, and nuclear weapons), New advancements in Chemistry, Noble laureates in Chemistry. Peaceful uses of chemistry, International Regulation of Biological and Chemical or Weapons of Mass Destruction.

Student will contact the respective mentor/seminar coordinator at allocated schedule to:
l. Conduct the literature survey of the topic allotted.
2. In the next step the student will prepare a detail report in consultation with mentor.
3. The student will learn from the mentor how to prepare presentations.
4. The student will give presentations before the mentor at allotted time schedules regularly.
5. Final seminar of students will presented before the committee consisting of all faculty members of Chemistry and submit their reports duly signed by mentors on the dates notified to them.

SEMESTER III

Semester	**III**
Course Code	**CHM601**
Course Title	**Spectroscopy-2(Techniques for Structural elucidation of Inorganic Compounds)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	To impart the knowledge of principles of electronic, rotation, vibration, laser, NMR, FTIR spectroscopy and their applications for structure elucidation of inorganic compounds.
Course Outcomes	The students will be able to: 1. Understand the basic concepts and principles of rotational and vibrational spectroscopic methods. 2. Apply various spectroscopic methods for structure elucidation of different inorganic compounds. 3. Comprehend the basic knowledge of X-ray spectroscopy and physical techniques for analysis of different medical diagnostics.

UNIT I:

General Introduction to spectroscopy: Nature of radiation, energies corresponding to various kind of radiation, energies for atomic and molecular transitions.

Electronic Spectroscopy: Electronic transition, energy of electronic transition, selection rules, the Franck-Condon principle.

Microwave Spectroscopy: Classification of molecules, rigid rotor model, effect of isotopic substitution on the transition frequencies, intensities of spectral lines, non-rigid rotor, Stark effect, applications.

Raman Spectroscopy - Classical and quantum theories of Raman Effect, pure rotational, vibrational and vibrational-rotational Raman spectra, selection rules, mutual exclusion principle, resonance Raman Spectroscopy, surface enhanced Raman spectroscopy, coherent anti stokes Raman spectroscopy.

UNIT II:

Vibrational /IR Spectroscopy: Theory of IR absorption, types of vibrations, observed number of modes of vibrations, Intensity of absorption bands, theoretical group frequencies. Review of harmonic oscillator, Selection rules, vibrational energies of diatomic molecules, zero point energy, force constant and bond strength, anharmonicity, vibration-rotation spectroscopy, Morse potential energy diagram, P, Q, R branches, vibrations of polyatomic molecules, overtones, hot bands and applications. Factors affecting group frequencies and band shapes (Physical state,

vibrational coupling, electrical effects, resonance, Inductive effects, Ring strain) vibrational-rotational fine-structure. Experimental method.

Application of IR to the following: Distinction between: Ionic and coordinate anions such as NO_3, SO_4 and SCN b) Distinction between : Lattice and coordinated water. Mode of bonding of ligands such as urea, dimethylsulphoxide and hexamethylphosphoramide.

UNIT III

Nuclear Magnetic Resonance (NMR) and Electron Spin Resonance (ESR) Spectroscopy: Basic concepts of NMR with emphasis on 31P, 19F, 29Si, 11B, 10B, 57Se, 125Te, 95Mo, 109Ag, 195Pt, 119 Sn and an explanations with appropriate examples. Historical introduction to magnetic resonances, chemical shift, mechanism of electron shielding and factors contributing to the magnitude of chemical shift, Nuclear overhausser effect, Lanthanide shift reagents and NMR spectra of paramagnetic complexes. Experimental technique. NMR in medical diagnostics. NMR study in Fluxional organometallic compounds**: T**rigonal bipyramidal molecules, Systems with coordination number six or more and organometallic molecules (PF_5, $Ti(acac)_2Cl_2$, $Ti(acac)_2Br_2$, $Ta_2(OMe)_{10}$,).

ESR: Hyperfine coupling, spin polarization for atoms and transition metal ions, spin orbit coupling and significance of *g*-tensors, application of transition metal complexes (having one unpaired electron) including biological systems.

NQR: Nuclear Quadrupole Resonance Spectroscopy: Introduction, effects of magnetic field on the spectra. Relationship between the electric field gradient and molecular structure. Interpretation of eQ, data, the effect of crystal lattiee on the magnitude of eQ4, double resonance technique,

UNIT IV

Mossbauer Spectroscopy:- Basic concepts (Radiation Source, Mossbauer Nuclei, Recoiless gamma resonance fluorescence. Use of the Doppler effect of vary the γ-ray energy). The effect of Isomeric shift, quadrupole hyperfine interaction and magnetic hyperfine interaction on MB spectra, MB experiment, Application of MB spectroscopy in structural determination of the following:

i) High spin Fe (II) and Fe (III) halides FeF_2, $FeCl_2.2H_2O$, FeF_3, $FeCl_3.6H_2O$. Low spin Fe(II) and Fe(III) Complexes-Ferrocyanides, Ferricyanides, Prussian Blue.
ii) Iron carbonyls. $Fe(CO)_5$, $Fe_2(CO)_9$ and $Fe_3(CO)_{12}$
iii) Inorganic Sn(II) and Sn(IV) halides.

Lasers and Laser Spectroscopy: Principles of laser action, pulsed lasers, examples of lasers: He-Ne, Nd-YAG, dye lasers.

Photoelectron spectroscopy: The photoelectric effect, UV photoelectron spectroscopy UPES, X-ray photoelectron spectroscopy XPES.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Physical methods in Inorganic Chemistry	R.S.Drago.	Reinhold Publishing Company (1965).
2	Modern Optical methods of Analysis	Eugens D.Olsen	McGraw-Hill Companies

3	Infrared spectra of Inorganic and coordination compounds	Kazuo Nakamoto	Wiley-Interscience; 6 edition
4	Basic Inorganic Chemistry	F.A.Cotton	John Wiley & Sons, 2009
5	Spectroscopy in Inorganic Chemistry	Rao & Ferraro	Academic Press, Inc. (1971)
6	A New Concise Inorganic Chemistry	J.D.Lee	John Wiley and Sons Ltd

Semester	**III**
Course Code	**CHM603**
Course Title	**Electrochemistry & Surface Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as one of the main Subject
Course Objective	This study Deals with the properties of surfaces or phase boundaries and with the chemical changes occurring at a surface or interface.
Course Outcomes	The students will be able to: 1. Acquire basic knowledge of Electrochemistry of electrode electrolyte interface and properties of surfaces or phase boundaries. 2. Understading basic concepts of electro chemistry, redox processes in electrochemical systems, EMF, pH and their applications 3. Apply Debye-Huckel theory, activity and activity coefficient ad Application of homogeneous and heterogeneous catalysis in chemical synthesis

Unit-I

Adsorption:

Surface tension, capillary action, pressure difference across curved surface (Laplace equations), vapor pressure of droplets (Kelvin equation), Gibbs adsorption isotherm, estimation of surface area (BET equation), surface films on liquids (Electro-kinetic phenomena), catalytic activity at surfaces.

Adsorption and Catalysis: Colloids and their stability, Adsorption of solids, Gibbs adsorption isotherm, BET adsorption isotherm, Langmuir and Fredulich Isotherms. Homogeneous catalysis and heterogeneous catalysis, enzyme catalysis. Michealis-Menten mechanism, Lineweaver-Burk Plot, competitive, non-competitive and uncompetitive bindings, kinetics of catalytic reactions.

Unit-II

Electrochemistry: Oxidation numbers. Redox potential. Electrochemical series. Redox indicators. Electrochemical cell reactions, Nernst equation, Electrode Kinetics, electrical double layer, electode/electrolyte interface, Batteries, primary & secondary Fuel Cells, Introduction to corrosion, homogeneous theory, forms of corrosion, corrosion monitoring and Prevention corrosion and corrosion prevention. Activity-coefficients, mean activity coefficients; Debye-Huckel treatment of dilute electrolyte solutions, derivation of Debye-Huckel limiting law, extended Debye-Huckel law and conductometric titrations.

Unit-III

Electrochemical Cells: Concentration cells with and without liquid junction, thermodynamics of reversible electrodes and reversible cells, potentiometric titration.

Colloidal State

Classification of colloids, charge and stability of colloidal dispersions, Hardy-Schulze Law, gold number, electrical properties of colloids, electrical double layer and its structure, Stern's theory of double layer, zeta-potential, electrophoresis and electro-osmosis, emulsions and their classification, emulsifiers, gels and their classification, thixotropy. Association colloids; miceller formation, cmc, soap action, Application of colloids.

Micelles: Surface active agents, classification of surface active agents, micellization, ydrophobic interactions, critical micellar concentration (CMC), factors affecting CMC of surfactants, counter ion binding to micelles, thermodynamics of micellization – phase separation and mass action models, solubilization, micro emulsion, reverse micelles.

Unit-IV

Voltmametry and Polarography: Polarography, polarographic cells, polarogram, interpretation of polarographic waves, equation for the polarographic waves, effect of complex formation on polarographic wave, polarograms for irreversible reactions, dropping mercury electrode, current variations during life time of a drop, merits and demerits of dme, polarographic diffusion currents, Ilkovic equation, capillary characteristics, temperature, polarograms for mixture of reactants, anodic and cathodic waves, factors affecting polarographic currents, applications of polarography, treatment of data, organic and inorganic polarographic analysis, voltammetry at solid electrodes, cyclic voltammetry and interpretation of data, , pilot-ion and standard addition method for quantitative analysis

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Surface Chemistry	E.M. Mc Cash	Oxford University Press, Oxford (2001).
2	An Introduction to Liquid State	P.A. Eglestaff	Academic Press.
3	Electrochemical methods, Fundamentals and Methods	A.J. Bard, L.R. Faulkner,	Wiley, 1980.
4	Physical Electrochemistry- Fundamentals, Techniques and Applications	EliezerGileadi,	Wiley-VCH 2011.

Semester	**III**
Course Code	**CHM605**
Course Title	**Advanced Coordination Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	The main goal of this subject is to study the coordination complexes, reactions and their applications.
Course Outcomes	The students will be able to: 1. Understand Formation, reaction mechanism of coordination complexes, their Kinetic and thermal stability, and determinations. 2. stability of coordination complexes. 3. Able to interpret the electronic and magnetic properties of coordination compounds.

Syllabus

Unit-I

Metal-Ligand Equilibria in Solution

Stepwise and overall formation constant and their interaction, trends in stepwise constants,f actors affecting the stability of metal complexes with reference to the nature of metal ionand ligand, chelate effect and its thermodynamic origin, determination of binary formation constants by spectrophotometry and potentiometric (pH) methods.

Unit-II

Theoretical and Practical aspects of Magnetism in Coordination Complexes Determination of state functions of R-S terms of d2 and p2, transition metal ions.

Derivation of Van Vleck's expression and mS+O formula, Quantization of orbital contribution in d1 ion and quenching in cubic crystal field.

Magnetic moments based on crystal field ground term, Perturbation Theory and its application, Spin orbit coupling operator for magnetic susceptibility and magnetic moment of T terms and A, E terms .

Anomalous magnetic moments in magnetically dilute and concentrated system in various symmetrical Environments of coordination complexes.

Unit-III

Reaction Mechanisms of Transition Metal Complexes Introduction, potential energy diagram and reactivity of metal complexes, ligand substitution reactions, labile and inert metal complexes, acid hydrolysis, factors affecting acid hydrolysis, base hydrolysis, conjugate base mechanism, anation reaction, substitution reactions in square planar complexes, trans effect, mechanism of the substitution reaction reactions without metal ligand bond cleavage, electron transfer processes outer and inner sphere.

Reactions of coordinated ligands Non-chelate forming reactions: Reaction of donor atoms(Halogenation of coordinated N atoms, Alkylation of coordinated S and N atoms, Solvolysis of coordinated phosphorus atoms). Reactions of non-donor atoms (nucleophilic behaviour of the ligand, electrophilic bahaviour of the ligand).

Chelate ring forming reactions: (reactions predominantly involving thermodynamic template effects, reactions predominantly involving kinetic affects).:Chelate modifying reactions

Unit-IV

Ligand field theory and molecular orbital theory; nephelauxetic series, structural distortion and lowering of symmetry, electronic, steric and Jahn-Teller effects on energy levels, conformation of chelate ring, structural equilibrium, magnetic properties of transition metal ions and free ions presentive, effects of L-S coupling on magnetic properties, quenching of orbital angular momentum by crystal fields in complexes in terms of splitting. effect of spin-orbit coupling and A, E and T states mixing.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Magnetism and Transition Metal Complexes	F. E. Mabbs and D. J. Machin	(Chapman and Hall) London (1973).
2	Introduction to Magnetochemistry	A. Earnshaw	Academic Press, (1968)
3	Elements of Magnetochemistry	R. L. Dutta and A. Syamal	Affiliated East/West Press Pvt. Ltd. 2007.
4	Inorganic Chemistry	D.F.Shriver, P.W.Atkins and C.H.Langford,	Oxford, 2nd. edn. 1994.
5	An Introduction to Inorganic Chemistry	K.F.Purcell and J.C.Kotz	Saunders 1990, Chapter 14.
6	Organotransition Metal Chemistry	Anthony F.Hill	Royal Society of Chemistry, Tutorial Chemistry Text, 2002. Chapters 1 to 7.
7	Organometallics: A concise Introduction	Ch.Elshebroicn and A Salzer	VCH, Chapters 12 to 16.
8	Comprehensive Coordination Chemistry	Vol.1. G Wilkinson (Ed)	Wiley, New York, 1967
9	Organotransition Metal Chemistry: Applications to Organic Synthesis	S.G.Davies	Pergamon 1982

Semester	**III**
Course Code	**CHM607**
Course Title	**Environmental Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To acquire the knowledge of different chemical phenomena as applied to environmental interfaces, policies as guidelines emanating from these phenomena and water/wastewater treatment techniques.
Course Outcomes	The students will be able to: 1. Comprehend Basic chemical processes in the air water and soil environment 2. Understand & Propose policies as guidelines regarding different environmental interfaces. 3. Apply different chemical phenomena as applied to environmental interfaces.

Unit-I

Environment: Introduction. Composition of atmosphere, vertical temperature, heat budget of the earth atmospheric system, vertical stability atmosphere. Biogeochemical cycles of C, N, P, S and O. Bio-distribution of elements.

Aquatic chemistry: Surface, ground water, marine and brackish water resources - assessment and utilization; Rivers and Lakes in India; hydrological cycle; Structure and properties of water, Aquatic pollution - inorganic, organic, pesticide, agricultural, industrial and sewage, detergents, oil spills and oil pollutants. Water quality parameters - dissolved oxygen, biochemical oxygen demand, solids, metals, content of chloride, sulphate, phosphate, nitrate and micro-organisms. Water quality standards, Analytical methods for measuring BOD, DO, COD, F, Oils, residual chloride and chlorine demand, Purification and treatment of water. Environmental Issues: Ground water depletion; Water logging and salinity; Water Conservation and management techniques; Rain water harvesting; Watershed management; Eutrophication; Restoration of Lakes, transboundary river water sharing and interlinking of rivers. Environmental chemistry of arsenic, chromium.

Unit-II

Water treatment Technologies: Chemical and Physical Methods of wastewater treatment with emphasis on sedimentation, coagulation, adsorption, water softening, defluoridation and ion exchange process.

Membrane Processes: Reverse Osmosis, Types of membrane, characterization of membranes, nano-membranes and their formation, efficiency of different membranes in removal of different elements.

Biological wastewater treatment including Activated sludge process, trickling filter and Membrane bioreactor, biological treatment processes - process description, design and application. Chemical potential, fugacity and its application to fugacity model.

Unit-III

Chemistry of Soil: Physio-chemical composition of soil, humus, inorganic and organic components of soil, nutrients (NPK) in soil, significance of C:N ratio, cation exchange capacity (CEC), reactions in soil solution, ion exchange (physiosorption), ligand exchange (chemisorption), complexations, chelation; precipitation / dissolution.

Environmental Geochemistry: Concept of major, trace and REE. classification of trace elements, mobility of trace elements, geochemical cycles.

Waste Management: Biomass waste management, biomedical waste management and chemical waste management, design and construction of waste management site. Regulations for waste management.

Unit-IV

Atmospheric chemistry: Composition of air, Chemical speciation, particles, ion and radicals, Formation of particulate matter, Photochemical reactions in the atmosphere, Chemistry of air pollutants, Photochemical smog, Acid rain, Analytical methods for measuring air pollutants. Continuous monitoring . Ozone Chemistry and Montreal Protocol, Greenhouse gases and Global warming, Clean Development Mechanism and Kyoto Protocol, Persistant Organic Pollutants (POP) and Stockholm Convention.

Sources of Natural and Artificial Radiations: Dosimetry, types of dosimeters, radioactive substances, applications and handling of isotopes and other radionuclides in environment.

Biochemical and Toxicological aspects of arsenic, cadmium, lead, mercury, carbon monoxide, O3, PAN, MIC and other carcinogens.

Environmental Toxicology: Chemical solutions to environmental problems, biodegradability, principles of decomposition, better industrial processes.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Environmental Chemistry	S. E. Manahan	Lewis Publishers
2	Environmental Chemistry	Sharma & Kaur	Krishna Publishers
3	Environmental Chemistry	A. K. De	Wiley Eastern
4	Environmental Pollution Analysis	S.M. Khopkar	Wiley Eastern
5	Standard Method of Chemical Analysis	F.J. Welcher	Vol. III, Van Nostrand Reinhold Co
6	Environmental Toxicology	Ed. J. Rose	Gordon and Breach Science Publication
7	Elemental Analysis of Airborne Particles	Ed. S. Landsberger and M. Creatchman	Gordon and Breach Science Publication
8	Environmental Chemistry	C. Baird, W. H. Freema	Wiley Eastern

Semester	**III**
Course Code	**CHM609**
Course Title	**Recent Trends in Inorganic Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To study recent advancements in Inorganic Chemistry with special emphasis to Inorganic Photochemistry, advanced catalysis.
Course Outcomes	The students will be able to: 1. Understand electronic structure of a variety of d orbital metal complexes and recent advancements in Inorganic Chemistry. 2. Acquire knowledge of Inorganic Photochemistry, Oxidative-Addition and Migration reactions 3. Use of Transition Metal Compounds with Bonds to Hydrogen and advanced catalysis.

Unit-I

Photo Inorganic Chemistry: Basics of photochemistry- Absorption, excitation, photochemical laws, quantum yield, electronically excited states, life times- measurements of the times Flash photolysis, energy adiddipation by radiative and non-radiative processes, absorption spectra, Franck-Condon principle, photochemical stages-primary and secondary processes, Kashia's rule, Thexi state, Photo substitution reactions, Adamson's rules, Photo substitution reactions of Cr(III)-Polypyridyls, Rh(III) Ammine Complexes, Ru-Polypyridyl complexes, Ligand photo reactions, photoredox reactions, comparison of Fe(II) and Ru(II) complexes, Photo reactions and Solar energy conversions, Photo synthesis in plants and Bacterio chlorophyll photosynthesis, photolysis of water using Inorganic precursors.

Unit-II

Oxidative-Addition and Migration (Insertion Reactions): Introduction: Acid base behaviour of metal atoms in complexes, Protonation and Lewis Base behaviour, acceptor properties of Lewis acidity of complexes, oxidative addition and reductive elimination, addition of specific molecules, Hydrogen addition, HX additions, Organic halides addition of some other molecules productive elimination, migration (Insertion) reaction promotion of alkyl migration, insertion of CO into M-H bonds, other aspects of CO insertion reactions, transfer of other molecules, CO_2, SO_2 , NO_2 , RCM, Insertion of alkenes and C-C unsaturated compounds, Cleavage of C-H bonds; alkane activation, Cyclometallation reactions. Reactions of free hydrocarbons.

Unit-III

Transition Metal Compounds with Bonds to Hydrogen : Characteristics of hydride complexes, synthetic methods, chemical behaviour of hydride compounds, mononuclear polyhydrides, homoleptic polyhydride anions; carbonyl hydrides and onion. Molecular hydrogen

compounds; metal hydrogen interaction with C-H bonds; MH interactions; complexes of boron hydride and aluminohydrides, synthetic applications of metal hydrides.

Unit-IV

Transition Metal Complexes in Catalysis : Hydroformylation of unsaturated compounds, Reductive carbonylation of alcohols and other compounds; Carbonylation Reaction: Methanol and methyl acetate, Adipic ester. Synthesis and other carbonylation reactions, decarbonylation reactions. Catalytic addition of molecules to C-C multiple bonds homogeneous hydrogenation, hydrocyanation of unsaturated compounds, hydrosilation of unsaturated compounds, hydrocyanation of alkenes, Polymerization, Oligomerisation and metathesis reactions of alkenes and alkynes, Ziegler-Natta polymerization of ethylene and propylene oligomerisation and related reactions, Cluster compounds in catalysis, supported homogeneous and phase transfer catalysis, Oxidation reaction: Oxidative carbonylations, Palladium catalysed oxidation of ethylene, Acrylonitrile synthesis, oxygen transfer from peroxo- and oxo- species, oxygen transfer from NO2 groups.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Concepts of Inorganic Photochemistry	A. W. Adamson and P. D. Fleischauer	Wiley.
2	Inorganic Chemistry	W.W. Porterfield	A Unified Approach
3	Advanced Inorganic Chemistry, 5th ed	F.A. Cotton and G. Willkinson	John Wiley & Sons, New York.
4	Organometalics: A Concise Introduction, 2nd Ed.	C. Elschenbroich and A. Salzer	VCH 1992.

Semester	**III**
Course Code	**CHM615**
Course Title	**Pharmaceutical Chemistry & Drug Design**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge about process of drug discovery & drug design
Course Outcomes	The students will be able to: 1. Acquire knowledge of Basic process of drug discovery & drug design. 2. Understanding of drug-receptor interactions and various drug mechanism. 3. Prediction of ligand interactions with the active site of receptor in novel drug design and discovery.

Unit-1

Drug Discovery and Drug Development: Introduction, Present and Past, Drugs and the medicinal chemist, Classification of drugs, Drug targets specification, Choice of Bioassay, In Vivo and in Vitro tests, Pit falls.

Drug Action at Receptors: Receptor role, Neuro-transmitters and Hormones, Change of shape by the receptors, Design of Agonists and Antagonists, Drug action on DNA and RNA.

Drug administration: Introduction, oral administration, sublingual administration, rectal administration, epithelial administration, inhalation, injection and implants.

Unit-II

Pharmacokinetics: Drug distribution and survival, Pharmacokinetic issues in drug design like Chemical and Metabolic stability, Hydrophilic / hydrophobic balance, Ionization, size and number of hydrogen bonding interactions, Drug dose levels, solubility and membrane permeability, variation of different groups to alter polarity.

Prodrugs: Introduction, Effect of prodrugs on: improved membrane permeability, prolonged drug activity, masking drug toxicity and side effects, increased chemical stability, targeting of drugs, prodrugs activation by external influence.

Unit-III

Drug Design, A Rational Approach: Introduction-analogues and prodrugs – concept of "lead" ,Quantum mechanical approach, Molecular orbital approach, Molecular connectivity approach, General considerations-tailoring of drugs.

Drug Design, Drug-Target Interactions: Introduction, Variation of Substituent, Expansion of the Structure, Chain expansion/Contractions, Ring expansion/Contractions, Ring Variation, Ring Fusions, Isosteres. Screening of natural products-Isolation and purification-structure determination, structure activity relationships

Unit-IV

Quantitative Structure-Activity Relationships (QSAR): Introduction, Hydrophobicity, Electronic effects, Steric factors, physicochemical parameters

Structural Features and Pharmacological Activity: The influence of steric factors, optical, geometrical isomerism, conformational isomerism and pharmacological activity.

Combinatorial Synthesis-The design of compound libraries and their application to drug discovery: application, combinatorial chemistry, future development and lead optimization, design based on structural information.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Textbook of Pharmacology	W. C. Bowman, and M. J. Rand	Blackwell Scientific 1980
2	Medicinal Chemistry-the role of organic chemistry in drug research	C. R. Ganellin, and S. M. Roberts	Academic Press 1993.
3	Medicinal Chemistry-principles and practice	F. D. King	The Royal Society of Chemistry 1994.
4	Burger's Medicinal Chemistry and drug discovery	M. E. Wolff	5 th edition Volume 1-5. Wiley 1995
5	The Organic Chemistry of Drug Design and Drug Action,	R. B. Silverman	Academic Press Inc. London2nd Edition, 2004
6	An introduction to Medicinal Chemistry	Graham L. Patrik	Oxford University Press, 3rd edition, 2005
7	Pharmaceutical Process Validation	I. R. Berry & R. A. Nash	Academic Press, London, 3rd Edition, 2003

Semester	**III**
Course Code	**CHM617**
Course Title	**Bio-Organic Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To provide knowledge of structure, function, and physicochemical properties of biomolecules. To aware students about the metalloenzymes, heme proteins, oxygen carriers, and non-heme proteins and therapatic Agents.
Course Outcomes	The students will be able to: 1. Understand structure, function and physicochemical properties of biomolecules. 2. Interpret Structure & Properties of enzymes, Mechanism of Enzyme Action metalloenzymes heme proteins and oxygen carriers. 3. Apply and use of non-heme proteins and therapatic Agents.

Unit-I

Enzymes : Basic considerations. Proximity effects and molecular adaptation. Introduction and historical prospective, chemical and biological catalysis, remarkable properties of enzymes like catalytic power, specificity and regulation. Nomenclature and classification, extraction and purification. Fischer's lock and key and koshland's induced fit hypothesis, concept and identification of active site by the use of inhibitors, affinity labelling and enzyme modification by site-directed mutagenesis.

Co-Enzyme Chemistry: Cofactors as derived from vitamins, coenzymes, prosthetic groups, apoenzymes.Structure and biological function of coenzyme A, thiamine pyrophosphate, pyridoxal phosphate, NAD+, NADP+, FMN, FAD, LIPOIC ACID, vitamin B12. Mechanisms of reactions catalysed by the above cofactors.

Unit-II

Mechanism of Enzyme Action : Enzyme kinetics, Michaelis-menten and lineweaver-Burk plots, reversible andirreversible inhibition. Transition-state theory, orientation and steric effect, acid-base catalysis, covalent catalysis, strain or distortion. Examples of some typical enzyme mechanisms for chymotrypsin, ribonucleases, lysozyme and carboxypeptidase A
Metalloenzymes: Definitions: Apoenzyme, Coenzyme, Metalloenzyme, structure and functions of carbonic anhydrase A & B, carboxy peptidases.

Unit-III

Vitamins: Introduction of fat soluble and water soluble vitamins, sources, structure, requirements and functions of vitamin A,D,E and vitamin B1 and C.
Antibiotics: ß- Lactumrings, structure and synthesis of penicillin – G, penicillin-V, Amoxycillin, cholamphenicol, streptomycin.

Unit-IV

Transport and storage of metals: The transport mechanism, transport of alkali and alkaline earth metals, ionophores, transport by neutral macrocycles and anionic carriers, sodium/potassium pump, transport and storage of Iron (Transferrin & Ferritin). Transport of Iron in microorganisms (sidrophores), types of sidrophores (catecholate and Hydroxmato siderophores).

Inorganic compounds as therapatic Agents:- Introduction chelation therapy, synthetic metal chelates as antimicrobial agents, antiarthritis drugs, antitumor, anticancer drugs (Platinum complexes), Lithium and mental health.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	The Inorganic Chemistry of Biological processes	M. N. Hughes	John Wiley & Sons Ltd
2	Medicinal Chemistry-the role of organic chemistry in drug research	C. R. Ganellin, and S. M. Roberts	Mount Kisko, NY 1973
3	Bio Inorganic Chemistry	Robert Wittay	
4	Advanced Inorganic Chemistry (4^{th} Edn)	Cotton and Wilkinson	
5	Topics in current chemistry (Inorganic Biochemistry) vol. 64 (1976)	Davison and Coworkers	
6	Inorganic chemistry	James E. Huheey.	

Semester	**III**
Course Code	**CHM619**
Course Title	**Material Chemistry**
Type of course	Theory Course
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge of materials, their characteristics and physical functions
Course Outcomes	The students will be able to: 1. Gain knowledge of Advanced Materials, their characteristics and physical functions. 2. Acquire knowledge of different types of materials like Glasses, Ceramics, Composites and Nanomaterials. 3. Apply Materials for Solid State Devices and Molecular Conductor.

UNIT-I

Magnetic Materials (Ferrites) Introduction, structure and classification, hard and soft ferrites, synthesis of ferrites by various methods (precursor and combustion method), characterization of ferrites by Mossbauer spectroscopy, significance of hysteresis loop and saturation magnetization in ferrites, magnetic properties of ferrites, applications of ferrites.

Glasses, Ceramics, Composites and Nanomaterials

Glassy state, glass formers and glass modifiers, applications. ceramic structures, mechanical properties, clay products. microscopic composites; dispersion-strengthened and particle-reinforced, fibre-reinforced composites, macroscopic composites, nanocrystalline phase, preparation procedures, special properties, applications.

Unit II

Mesmorphic behaviour, thermotropic liquid crystals, positional order, bond orientational order, nematic and smectic mesophases; smectic - nematic transition and clearing temperature - homeotropic, planar and sCHMieren textures, twisted nematics, chiral nematics, molecular arrangement in smectic A and smectic C phases, optical properties of liquid crystals. dielectric susceptibility and dielectric constants. lyotropic phases and their description of ordering in liquid crystals.

Materials for Solid State Devices

Rectifiers, transistors, capacitors –IV-V compounds, low–dimensional quantum structure; optical properties.

Unit III

Types of ionic conductors, mechanism of ionic conduction, interstitial jumps (Frenkel); vacancy mechanism, diffusion superionic conductors; phase transitions and mechanism of conduction in superionic conductors, examples and applications of ionic conductors.

Molecular Conductor: Oligo (phenylene vinylene)s, oligo(phenylene ethynylene)s, oligo (eneyne)s, oligo(thiophene vinylene), oligo (thiphene ethynylene) etc. and their applications.

Unit IV

Preparation and characterization of silica and zirconia based stationary phases by (a) dynamic chemical modification, in which chiral selector is adsorbed on the surface of the zirconia by physical forces, (b) permanent chemical modification, in which a CS is chemically bonded onto the zirconia surface, and (c) physical screening, in which zirconia surface is coated with a polymer or carbon layer, and their application in chiral separations by LC

Fullerenes, Carbon Nanotubes and Graphene: Types and Properties, methods of preparation and separation of carbon nanotubes, applications of fullerenes, CNTs and graphene.

Text and Reference Books:

S.No.	Name/Title	Author	Publisher
1	Fundamentals of Materials Science and Engineering: An Integrated Approach.	Callister Jr, W. D., and Rethwisch, D. G.	John Wiley and Sons. (2012).
2	Materials Science for Engineers.	Anderson, J. C., Leaver, K. D., Rawlings, R. D., and Leevers, P. S.	CRC Press. (2004).
3	Principles of the Solid State.	Keer, H. V.	New Age International. (1993

Semester	**III**
Course Code	**CHM611**
Course Title	**Inorganic Chemistry Practical-III**
Type of course	Practical Course
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To aware students about
Course Outcomes	The students will be able to: 1. Acquire knowledge of basic reparation routes of inorganic compounds. 2. Apply semi-micro qualitative analysis of mixtures & gravimetric analysis for different cations and anions. 3. Apply different types of Potentiometry and pHmetry titrations.

Inorganic Preparations

Preparation of mercury tetraisothiocyanatocobaltate(II). Determination of its magnetic moment and interpretation of its IR spectrum.

Preparation of nitro-and nitrito-pentaamminecobalt(II) chlorides from chloropenta amine cobalt(III) chloride. Recording and interpreting their electronic and IR spectra.

Preparation and resolution of tris(ethylenediamine)cobalt(II) ion. Measurement of optical rotation of these resolved complexes.

Preparation of diaquotetraacetatedicopper(II). Determination of its magnetic susceptibility and interpretation of E.P.R., electronic absorption and IR spectra.

Preparation of hexaamminenickel(II) chloride and tris(ethylenediamine)nickel(II) chloride. Interpretation of their electronic absorption spectral data and calculation of β and 10Dq values. Measurement of magnetic susceptibility, calculation and interpretation of the values.

Preparation of lead tetraacetate.

Preparation of potassium trioxalatoaluminate(III) trihydrate. Its TGA and DTA studies and its interpretation of its IR data.

Preparation of sodium tetrathionate, potassium dithionate, and interpretation of their IR spectra.

13 Preparation of iron(II) oxalate and potassium trioxalateferrate(III). Interpretation of their magnetic data, E.p.r. and Mossbauer spectra.

Electrogravimetry Titrations

Determination of Copper and Lead in a given sample of Brass Electrogravimetrically .

Determine coulometrically the concentration of Nickel and Cobalt from a given mixture.

Polarography methods

The polarographic Determination of Copper and Zinc in the given sample of Brass.

Study the polarographic waves produced by dissolved oxygen.

Plot a polarogram for a mixture of Cd^{2+}, Zn^{2+}, and Mn^{2+}, ions.

Potentiometry and pHmetry

To determine the dissociation constant of a dibasic acid(malonic acid)

The potentiometric titration of a mixture of Chloride and Iodide with $AgNO_3$.

To determine the degree of hydrolysis of aniline hydrochloride and hence hydrolysis constant of the salt.

Titration of Phosphoric acid solution with NaOH using quinhydrone electrode.
The Potentiometric Determination of Solute Species in a Phosphate Mixture
The Potentiometric Titration of Copper with EDTA.
Apply stripping methods to determine the concentration of lead in tap water.

Amperometric titrations:

Amperomertic titration of lead solution with potassium dichromate.
Amperometric titration of potassium sulphate solution with Lead nitrate solution.
Amperometric titration of nickel in solution with dimethyl glyoxime.
Determine transport number of silver and nitrate ions by Hittorf's method.
To study complex formation between Fe(III) and salicylic acid and find out the formula of the complex spectrophotometrically.
To study the kinetics of hydrolysis of crystal violet spectrophotometrically.
Determination of nitrite in water spectrophotometrically.
Determination of molecular weight of polymers by Tirbiditymetery.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Vogel's quantitative analysis 6 Edn	Mendham, Denny	Pearson Education 2002
2	Synthesis and Technique in Inorganic chemistry	G. S.Girlomi; R.J. Angleci	3rd edn.; University Science Books.
3	Advanced Practical Inorganic Chemistry	Ayodha Singh	Campus Books 2002

Semester	**III**
Course Code	**CHM613**
Course Title	**Physical Chemistry Practical-III**
Type of course	Practical Course
L T P	0 0 4
Credits	2
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To aware students about adsorption, chemical kinetics, solutions and phase equilibria and their applications.
Course Outcomes	The students will be able to: 1. Understand experimental techniques for controlling chemical reactions. 2. Aly and measure various physical and chemical properties of materials. 3. Design & carry out scientific experiments and result interpretation.

Adsorption

To study surface tension - concentration relationship for solutions (Gibbs equation).

To verify Freundlich and Langmuir adsorption isotherms for adsorption of acetic acid on activated charcoal.

To determine the freezing point depression constant of camphor using naphthalene as solute. Hence determine the molecular weight of acetanilide by Rast's micro method.

Determination of molecular weight of non-volatile and non-electrolyte/electrolyte by cryoscopic method and to determine the activity coefficient of an electrolyte

Phase Equilibria

Determination of partition coefficient of benzoic acid between organic solvent and water.

Determination of partition coefficient of iodine between water and octanol and determination of equilibrium constant of tri-iodide.

Determination of congruent composition and temperature of a binary system (e.g., diphenylamine benzophenone system)

To construct phase diagram of 3-component system (CH3COOH + CHCl3 + H2O)

Chemical Kinetics

Determination of the effect of (a) Change of temperature (b) Change of concentration of reactants and catalyst and (c) Ionic strength of the media on the velocity constant of hydrolysis of an ester/ionic reactions.

Determination of the velocity constant of hydrolysis of an ester/ionic reaction in micellar media.

Determination of the rate constant for the oxidation of iodide ions by hydrogen peroxide studying the kinetics as an iodine clock reaction.

To study kinetics of inversion of cane sugar by optical rotation measurement.

Determination of energy of activation of S2O82– +I□□ SO42–+I2 reaction

Studies on the effect of variation of ionic strength on the rate of S2O82– +I□□ SO42– + I2 reaction Curve fitting using linear and non-linear (Activation thermodynamic parameter, equilibrium thermodynamic parameter) regression analysis using software.

Refratrometry

(i)Determination of refractive indices (RI) of given liquids and determination of the concentration from RI.
(ii)Determine the specific refraction, molar refraction and atomic parachor with the help of Abbe's refractometer
(iii) Determination of specific and molar refraction of a liquid by Abbe refractometer.
(iv) To determine the molar refracitivity of water, DMF, Dioxane and mixtures of water, DFM, water-Dioxane and verify the refractivity rule.
(v)Predict about the interactions between components of mixture by plotting graph between refractive index and mole fraction.
(vi)Determine the refraction equivalents of C, H, and Cl atoms.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Experimental Physical Chemistry	Arthur M. Halpern, George C. McBane	Freeman,2006.
2	Experiments in Physical Chemistry, 5th ed.,	Schoemaker et al.	MGH, 1989
3	Chemistry Experiments for Instrumental Methods	Sawyer, Heineman, Beebe	Wiley, 1984 .
4	Physical Chemistry Practical.	Maity S., and Ghosh, N.(	New Central Book Agency (P) Ltd. 2012).
5	Senior Practical Physical Chemistry.	Khosla, B.D., Garg, V.C., and Gulati A.R.	S. Chand and Sons. (2007).
6	Advanced Practical Physical Chemistry.	Yadav, J. B.	Krishna Prakashan Media. (2006).
7	Experiments in Physical Chemistry,	Ghosh, J.C.	Bharati Bhavan. (1990).

Semester	**III**
Course Code	**CHM629**
Course Title	**Project (Phase-I)**
Type of course	Practical Course
L T P	0 0 8
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	The project would develop scientific aptitude, reviewing of literature, critical thinking, hypothesis development, experiment planning, synopsis writing, problem presentation and way to solve the problem.
Course Outcomes	The students will be able to: 1. Analyze current literature research for research topic of his/her area of expertise. 2. Design a research problem and prepare synopsis. 3. Plan future experiments in the laboratory.

- Project supervisor would be allocated at the start of the semester and research project would be undertaken in discussion with the project supervisor.
- At the end of the semester the student has to prepare a project report as per the university guidelines.
- Upon submission of the project report, the projects would be evaluated based on a project presentation.

SEMESTER IV

Semester	**IV**
Course Code	**CHM602**
Course Title	**Chemistry of Natural Products & Heterocyclic Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge about classification, occurrence and biosynthesis of various natural products and synthesis of organic compounds containing N, O, and S like compounds.
Course Outcomes	The students will be able to: 1. Gain Coherent and advanced knowledge of various types of natural products, their biosynthesis 2. Analyse structure, synthesise knowledge to identify and to identify complex structure of natural products 3. Acquaint knowledge about terpenoids, vitamins, alkaloids and hormones.

Unit-I

Terpenoids and Carotenoids: Classification, nomenclature, occurrence, isolation, general methods of structure determination, isoprene rule. Structure determination, stereochemistry, biosynthesis and synthesis of the following representative molecules: Citral, Geraniol, α-Terpeneol, Menthol, Farnesol, Zingiberene, Santonin, Phytol, Abietic acid and β-Carotene.

Alkaloids: Definition, nomenclature and physiological action, occurrence, isolation, general methods of structure elucidation, degradation, classification based on nitrogen heterocyclic ring, role of alkaloids in plants. Structure, stereochemistry, synthesis and biosynthesis of Ephedrine, (+)- Coniine, Nicotine, Atropine, Quinine and Morphine.

Unit-II

Steroids: Occurrence, nomenclature, basic skeleton, Diel's hydrocarbon and stereochemistry. Isolation, structure determination and synthesis of Cholesterol, Bile acids, Androsterone, Testosterone, Estrone, Progestrone, Aldosterone. Biosynthesis of steroids

Plant Pigments: Occurrence, nomenclature and general methods of structure determination. Isolation and synthesis of Apigenin, Luteolin, Quercetin, Myrcetin, Quercetin-3-glucoside, Vitexin, Diadzein, Butein, Aureusin, Cyanidin-7-arabinoside, Cyanidin, Hirsutidin. Biosynthesis of flavonoids: Acetate pathway and Shikimic acid pathway.

Unit-III

Nomenclature of Heterocycles: Replacement and systematic nomenclature (Hantzsch-widman System) for monocyclic fused and bridged hetrocycles

Aromatic and Non aromatic Heterocycles: General chemical behaviour of aromatic heterocycles classification (structural type) criteria of aromaticity(bond length ring current and chemical shift in H NMR- Spectra empirical resonance energy delocalization energy and Dewar

resonance energy)**,** Principles of heterocyclic synthesis involving cyclization reactions and cycloaddition reaction.
Strain-bond angle and torsional strains and their conseqences in small ring heterocycles. Conformation of six-membered heterocycles with reference to molecular Geometry

Unit-IV

Small Ring Heterocycles: Three-membered and four-membered heterocyclic –synthesis and reactions of aziridines, oxiranes, thiiranes, azetidines, oxetanes and thietanes

Benzo-Fused Five-Memberd Heterocycles

Synthesis and reaction including medicinal applications of benzopyrroles, benzofurans and benzothiophenes

Reagents in Organic Synthesis: Use of the following reagents in organic synthesis and functional group transformations; Complex metal hydrides, Gilman's reagent, lithium dimethylcuprate, lithium disopropylamide (LDA) dicyclohexylcarbodimide. 1,3-Dithiane (reactivity umpolung), trimethylsilyl iodide, Woodward and prevost hydroxylation, osmium tetroxide, DDQ, selenium dioxide, phase transfer catalysts, crown ethers and Merrifield resin, Peterson's synthesis, Wilkinson's catalyst, Baker yeast.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Organic Chemistry, Vol. 2, 5th edition	Finar, I.L.	ELBS, 1975.
2	1. Chemistry, Biological and Pharmacological Properties of Medicinal Plants from the Americas	Hostettmann, Kurt; Gupta, M.P.; Marston, A.	Harwood Academic Publishers.
3	Introduction to Flavonoids	Aggarwal, O.P.	2. Harwood Academic Publishers.
4	3. Natural Products: Chemistry and Biological Significance,	Mann, J.; Davidson, R.S.; Hobbs, J.B.; Banthrope, D.V.; Harborne, J.B.	. Longman, Esse
5	Organic Chemistry	Jerry March	Wiley & Sons
6	Heterocyclic Chemistry	Acheson.	Wiley-Interscience; 3rd edition (March 11, 1985)
7	Advanced Organic Chemistry	F.R.Carey, R.J. Sunberg.	Wiley Publishers
8	Highlights of Organic Chemistry	W.J.L. Nobel	An Advanced Text Book
9	Organic Chemistry	Jerry March	Wiley & Sons

Semester	**IV**
Course Code	**CHM604**
Course Title	**Bio-Inorganic Chemistry**
Type of course	CR Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	Main aim to Study the role of metals, enzymes, photosystems in biology. To provide knowledge of structure, function, and physicochemical properties of biomolecules.
Course Outcomes	The students will be able to: 1. Gain Coherent and advanced knowledge of various types of metals, enzymes, photosystems in biology. 2. Acquaint knowledge about the role of electron Transfer in Biology. 3. Analyse structure, function, and physicochemical properties of biomolecules.

Unit-I

Role of Metal ions in Biological Systems: Functions of metal ions in biological systems. Transport of ions through cell membrane - Na+/K+ Pump.

Natural oxygen carriers: Structure and function of Hemoglobin, Myoglobin, Hemerythrin & Hemocyanin. Mechanism of dioxygen binding with heme proteins. oxygen Transport in human body (-perutz machanism), Nature of Iron-dioxygen linkage in Hemoglobin, Model system - Model Synthetic complexes of Iron and Cobalt as Oxygen carrier, Cyanide poisoning and its remedy. Non-heme protiens (Hemerythrin & Hemocyanin).

Synthetic oxygen carriers: Oxygen molecule and its reduction products, model compounds for oxygen carrier (Vaska's Iridium cjomplex, cobalt complexes with dimethyl glyoxime and schiff base ligands).

Unit-II

Metalloporphyrins: Porphyrins and their salient features, characteristic absorption spectrum of porphyrins, chlorophyll (structure and its role in photosynthesis).

Photosynthetic Pigments: Complexes of Porphyrin. Redox mechanism in Photosystems (PS-I and PS-II), Cleavage of Water in PS-II.

Biological Nitrogen Fixation: Enzymetic reduction of Nitrogen to Ammonia - Nitrogenase Structure and mechanism. Molybdenum Nitrogenase - Spectroscopic and other studies. Model Systems for Nitrogenase.

Unit-III

Electron Transfer in Biology: Structure and Function of Metallic Proteins in Electron Transport Process, Cytochrome, Iron-sulphur Proteins.

Metal Storage, Transport and Biomineralisation : Metal Storage and Transport Structure and Function of Feritin, Transfrsin and Siderophores. Biomineralisation.

Calcium in Biology: Biochemistry of calcium as hormonal messenger, muscle contraction blood clotting, neurotransmitter, metals in the regulation of biochemical events. Role of Calcium in

living systems. Transport and regulation of Calcium - pathways and mechanism, Calcium Channels and pumps. Calcium binding proteins - Intracellular, Mediated membrane and Extracellular.

Unit-IV

Metal - Nuclic Acid Interactions: Metal ions and Nucleic Acids Interactions (binding) - Types & suitable examples. Metal complexes (e.g., Cisplatin) interaction with DNA. (b) Metals in Medicine: Metal Deficiency and disease. Toxic effects of Metals. Metals used for diagnosis. Metals used in Chemotherapy with particular reference to Anticancer drugs.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Principles of Bioinorganic Chemistry,	S. J. Lippard and J. M. Berg,	University Science Books (1994)
2	Bioinorganic Chemistry	Bertini, H. B. Grey, S. J. Lippard and J. S. Valentine	Viva Books Pvt. Ltd., New Delhi (1998
3	Biological Inorganic Chemistry: Structure and Reactivity	Harry B. Gray, Edward I. Stiefel et al.,	University Science Books.
4	Biological Inorganic Chemistry: An Introduction	Robert Crichton	.Elsevier Science (2008)
5	Bioinorganic Chemistry: Inorganic Elements in the Chemistry of Life, An Introduction and Guide	W. Kaim and B. Schwederski	Wiley, New York (1995)
6	Inorganic Chemistry of Biological Processes, 2nd Ed.	M. N. Hughes	JohnWiley & Sons, New York (1981),

Semester	**IV**
Course Code	**CHM606**
Course Title	**Instrumental Methods of Analysis**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge of various analytical and instrumental methods for chemical characterization and analysis.
Course Outcomes	The students will be able to: 1. Understand Coherent and advanced knowledge of various analytical and instrumental methods for chemical characterization and analysis. 2. Cognitive skills to analyse and apply analytical instrumental techniques for identification,characterization of compounds. 3. Apply theoretical and practical skills of the hyphenated instruments for identification of compounds.

Unit-I

Data Analysis: Linear regression, covariance and correlation coefficient. Standard reference materials, criteria for selection of analytical method. Uncertainties, Errors, calibrations, Mean, Standard Deviation , Least square fit,

Atomic Absorption Spectroscopy: General principles, instrumental set up and analytical procedures and applications, fluorescence spectrometry, flame AAS, electrothermal AAS (ETAAS).

Unit-II

Thermo-Analytical Method: Theory, instrumental requirements and methodology for thermo gravimetric analysis (TGA), differential thermal analysis (DTA) and differential scanning calorimetry (DSC), applications

Chromatographic Methods:

Partition and distribution, principles of chromatography, plate and rate theory. retention time and retention factor, resolution and separation factor; general idea about adsorption, partition and column chromatography, paper and thin layer chromatography, gas chromatography (GC) and high performance liquid chromatography (HPLC) - instrumentation, methodology and applications. SFC LC, hyphenated techniques, LC-MS and LC-MS/MS. Ion exchange resins and extraction, Ion Chromatography, anion supressors and ion speciation analysis.

Unit-III

Potentiometry – General principles, reference electrodes, ion selective electrodes, ion selective electrode construction, membrane electrode, glass electrodes, liquid membrane electrodes, biosensors ISFET and MOSFETS.

Coulometry: Basic principles of electrogravimetry, ohmic potential, kinetic and concentration polarization, overpotential, constant current and constant potential coulometry. coulometric titrations and application.

Voltammetry: Principles, dropping mercury electrode (DME), polarography, half-wave potential, diffusion current and Illkovic equation, different wave forms–linear scan, square scan and triangular scan, cyclic voltammetry, voltammograms. Anion/cation striping voltametry and its applications.

Unit-IV

Electrochemical Techniques: Conductometry, pH metry, Karl Fischer titration, cyclic voltametry , Polarography

Modern Methods of Surfaces and Crystal Analysis: SEM, TEM, STM, AFM, XRD: Instrumntation an

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Instrumental Methods of Analysis,	Willard, Merritt, Dean and Settle	CBS Publisher and Distributors.,1986.
2	Thermal Analysis,	W. W. Wendlandt and L. W. Collins,	Dowden Hutechin and Ross
3	Basic Concepts of Analytical Chemistry	S. M. Khopkar ,	Wiley Eastern
4	Thermal Methods of Analysis, Principles, Application and Problems,	J. Haines,, R.S.; Hobbs, J.B.; Banthrope, D.V.; Harborne, J.B.	Blackie Academic and Professional, 1994. Longman, Esse
5	Chromatographic Methods	A. Braithwaite and F. J. Smith	5th edn. Blackie Academic and Professional, London, 1996
6	Principles of Instrumental Analysis	Skoog, Holder, Nieman	Fifth edition Thomson Books ,1998.

Semester	**IV**
Course Code	**CHM608**
Course Title	**Nano-Science& NanoChemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart knowledge of nano chemistry and nanomaterials. carbon nanotubes and their applications
Course Outcomes	The students will be able to: 1. Acquire knowledge of Nanotechnology, properties and applications of nanomaterials. 2. Cognitive skills to analyse the methodology and fabrication and characterization of nanomaterials 3. Apply use of carbon nanotubes based nanomaterials.

Unit-I

Nanochemistry Basics: Nanochemistry, self assembly, Self assembling materials, two dimensional assemblies, Mesoscale self assembly, coercing colloids.

Chemical Patterning, Lithography & Nanocontact Printing: Soft lithography, Dip pen nanolithography, Nanoplotters, Nanoblotters,

Unit-II

Nanomaterials: Nanoparticles: zero dimensional nanostructure, homogeneous and heterogeneous nucleation, metallic nanoparticles- synthesis and applications; nanowires and nanorods: one dimensional nanostructures, spontaneous growth, VLS, electro spinning, lithography; thin film: two dimensional nanostructure- preparation techniques; Langmuir-Blodgett (LB) film growth techniques, photolithography properties and applications.

Unit-III

Nanorod, Nanotube, Nanowire Self- Assembly: Templating nanowires, nanorods, Nanorod devices, Nanowire sensors, diodes & transistors.

Carbon nanostructures: Carbon molecules, clusters, carbon nanotubes and their applications.

Unit-IV

Organic Compounds and Polymers: Nanocrystals, polymers, Supromolecular structures.

Scope and opportunities: Nanoscale materials, nanotechnology enabled sensors, microelectronics, drug delivery, Bionanoinformation.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Nanochemistry, A Chemical approach to Nanomaterials	G. A. Ozin & Andre, C. Arsenault	Royal society of Chemists, 2005.

2	Introduction to Nanotechnology	C. P. Poole, Jr., F. J. Owens	Wiley interscience
3	Real world cases in green chemistry	M.C. Cann and M. E. Connelly	ACS Publications.
4	Policies for cleaner Technologies	T. Clayton	Earthscan
5	New Trends in Green Chemistry	V. K. Ahluwalia and M. Kidwai	Anamaya Publishers, New Delhi.

Semester	**IV**
Course Code	**CHM610**
Course Title	**Green Chemistry**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To study about the greener methods which are eco friendly to environment like microwave induced greener synthesis, using ionic liquids or greener solvents, grinding conditions and ultrasound assisted green synthesis.
Course Outcomes	The students will be able to: 1. Acquire Coherent knowledge of concepts and tools of green chemistry and their importance in sustainable development. 2. Utilize abundantly available precursors for the production of value added chemicals. 3. Adopt and design solvent free synthesis strategies,Microwave assisted and sonicator in organic synthesis.

Unit-I

Designing a Green Synthesis: Choice of Starting Materials, Choice of Reagents, Choice of Catalysts, Choice of Solvents.

Basic Principles of Green Chemistry: Prevention of Waste/By-Products, Maximum Incorporation of the Reactants into the Final Product, Prevention or Minimization of Hazardous Products, Designing Safer Chemicals, Energy Requirements for Synthesis, Selection of Appropriate Solvent, Selection of Starting Materials, Use of Protecting Groups, Use of Catalyst, Products Designed Should be Biodegradable, Designing of Manufacturing Plants, Strengthening of Analytical Techniques.

Unit-II

Green Chemistry in Day-to-Day Life: Dry Cleaning of Clothes, Versatile Bleaching Agent.

Green Reagent: Dimethylcarbonate, Polymer Supported Reagents.

Green Catalysts: Acid Catalysts, Oxidation Catalysts, Basic Catalysts, Polymer Supported Catalysts.

Phase Transfer Catalysis in Green Synthesis: Introduction, Applications of PTC in Organic Synthesis, Oxidation Using Hydrogen Peroxide Under PTC Condition, Crown Ethers.

Unit-III

Microwave Induced Green Synthesis: Introduction, Applications - Microwave Assisted Reactions in Water, Microwave Assisted Reactions in Organic Solvents, Microwave Solvent Free Reactions (Solid State Reactions).

Ultrasound Assisted Green Synthesis: Introduction, Applications of Ultrasound.

Biocatalysts in Organic Synthesis: Introduction, Biochemical (Microbial) Oxidations, Biochemical (Microbial) Reductions, Enzymes Catalysed Hydrolytic Processes.

Unit-IV

Aqueous Phase Reactions: Introduction, Diels-Alder Reaction, Claisen Rearrangement, Wittig-Homer Reaction, Michael Reaction, Aldol Condensation, Knoevenagel Reaction, Pinacol Coupling, Benzoin Condensation, Claisen-Schmidt Condensation, Heck Reaction, Strecker Synthesis, Wurtz Reaction, Oxidations, Reductions, Polymerisation Reactions, Photochemical Reactions, Electrochemical Synthesis, Miscellaneous Reactions in Aqueous Phase.

Organic Synthesis in Solid State: Introduction, Solid Phase Organic Synthesis Without Using Any Solvent, Solid Supported Organic Synthesis.

Versatile Ionic Liquids as Green Solvents: Green Solvents, Reactions in Acidic Ionic, Liquids, Reactions in Neutral Ionic Liquids.

Unit-V

Synthesis Involving Basic Principles of Green Chemistry: Some Examples; Introduction, Synthesis of Styrene, Synthesis of Adipic Acid, Catechol and 3-dehydroshikirnic Acid (a potential replacement for BHT), Synthesis of Methyl Methacrylate, Synthesis of Urethane, An Environmentally Benign Synthesis of Aromatic Amines, Selective Alkylation of Active Methylene Group, Free Radical Brornination, Acetaldehyde, Furfural from Biomass, Synthesis of (S)-metolachlor, an Optically Active Herbicide, Synthesis of Ibuprofen, Synthesis of Paracetamol, Green Synthes~s Qf3-phenyl Catechol, Synthesis of Epoxystyrene, Synthesis of Citral, Synthesis of Nicotinic Acid, Use of Molting Accelerators to Replace More, Toxic and Harmful Insecticides, An Environmentally Safe Marine Antifoulant.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Green chemistry frontiers in benign chemical synthesis and processes	P. Anastas and H. Williamson	Oxford University Press.
2	Chemical management: Reducing wast and cost through innovative supply strategies	Lerma and W. Straat	Willey Sons
3	Real world cases in green chemistry	M.C. Cann and M. E. Connelly	ACS Publications.
4	Policies for cleaner Technologies	T. Clayton	Earthscan
5	New Trends in Green Chemistry	V. K. Ahluwalia and M. Kidwai	Anamaya Publishers, New Delhi.

Semester	**IV**
Course Code	**CHM612**
Course Title	**Industrial Chemical analysis & Quality Control**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To impart basic knowledge of basic Industrial Chemical analysis & Quality Control processes.
Course Outcomes	The students will be able to: 1. Acquire Coherent and advanced knowledge of the basic of Industrial Chemical analysis & Quality Control processes. 2. Analyze Chemical, biological and radiational hazards in laboratory and safetyfolled during analysis of Special Industrial Material. 3. Apply & Design analytical sample preparation and the analyse the clinical samples and chemical Sensors.

UNIT-I

Analytical Chemometrics : General introduction and its application in optimisation, Modelling and parameter estimation, Sampling, calibration, Factor analysis, Resolution, Signal processing, Structure-property relationship, Pattern recognition, Propagation of measurement uncertainties (inaccuracy and imprecision), Analytical validation techniques, Non-linear regression analysis, Good manufacturing practice (GMP), Good lab practice (GLP), lab and industrial safety.

UNIT-II

Analysis of Special Industrial Material (General Strategy for Analysis) : Analysis of dairy products, oils, soaps and synthetic detergents, food additives, petrochemicals (including liquid and gaseous fuels) pesticides, drugs and pharmaceuticals, fertilizers and paints.

UNIT-III

Clinical Analysis : Sampling and selective analysis of biological fluids (using routine and automatic instruments), glucose, bilirubins, total cholesterol, haemoglobin, creatinine, total proteins, albumin, urea-nitrogen, carticosteroids and barbiturates. Immunological methods of analysis: ELISA, RIA and Immunodiffusion.

UNIT-IV

Chemical Sensors : Principles, types of chemical sensors based on the modes of transductions, Types of chemical sensor based on the chemically sensitive materials (solid electrolyte, gas, semiconductor), Humidity sensors, Biosensors, Electrochemical sensors (Potentiometric sensors, Ion-selective electrodes, Membrane electrodes, Amperometric sensors, Clark and Enzyme electrodes).

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Green chemistry frontiers in benign chemical synthesis and processes	P. Anastas and H. Williamson	Oxford University Press.
2	Chemical management: Reducing wast and cost through innovative supply strategies	Lerma and W. Straat	Willey Sons
3	Real world cases in green chemistry	M.C. Cann and M. E. Connelly	ACS Publications.
4	Policies for cleaner Technologies	T. Clayton	Earthscan
5	New Trends in Green Chemistry	V. K. Ahluwalia and M. Kidwai	Anamaya Publishers, New Delhi.

Semester	**IV**
Course Code	**CHM614**
Course Title	**Polymer Science**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	Main aim to Study the polymers, their processing, structure, properties and analysis of polymers by various methods
Course Outcomes	The students will be able to: 1. Acquire Coherent knowledge of different polymers, their processing, structure, properties and mechanisms of polymerization. 2. Analyze number, weight and viscosity average molecular weights with various techniques 3. Apply & Design of methodologies for thermoplastic and thermosetting polymers, concept of conducting polymers and their applications.

Syllabus

Unit-I

Basics: Importance of polymers. Basic concepts: Monomers, repeat, units, degree of polymerization. Linear, branched and network polymers. Classification of polymers. Polymerization: condensation, addition, radical chain-ionic and co-ordination and copolymerization. Polymerization conditions and polymer reactions. Polymerization in homogeneous and heterogeneous systems.

Unit-II

Polymer Characterization: Polydispersion-average molecular weight concept. Number, weight and viscocity average molecular weights. Polydispersity and. molecular weight distribution. The practical significance of molecular weight. Measurement of molecular weights. End-group, viscocity, light scattering, osmotic and ultracentrifugation methods. Analysis and testing of polymerschemical analysis of polymers, spectroscopic methods, X-ray diffraction study. Microscopy. Thermal analysis and physical testing-tensile strength. Fatigue, impact. Tear resistance. Hardness and abrasion resistance.

Unit-III

Structure and Properties: Morphology and order in crystalline polymers-configurations of polymer chains. Crystal structures of polymers. Morphology of crystalline polymers, strain-induced morphology, crystallization and melting. Polymer structure and physical properties - crystalline melting point Tm - melting points of homogeneous series, effect of chain flexibility and other steric factors; entropy and heat of fusion. The glass transition temperature, Tg-Relationship between Tm and Tg, effects of molecular weight, diluents, chemical structure, chain topology, branching and cross linking. Property requirements and polymer utilization.

Unit-IV

Polymer Processing: Plastics, elastomers and fibres. Compounding. Processing techniques: Calendering, die casting, rotational casting, film casting, injection moulding, blow moulding, extrusion moulding, thermoforming, foaming, reinforcing and fibre spinning.

Unit-V

Properties of Commerical Polymers: Polyethylene, polyvinyl chloride, polyamides, polyesters, phenolic resins, epoxy resins and silicone polymers. Functional polymers - Fire retarding polymers and electrically conducting polymers. Biomedical polymers -contact lens, dental polymers, artificial heart, kidney, skin and blood cells.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Textbook of Polymer Science	F.W. Billmeyer Jr	Wiley
2	Polymer Science	V.A. Gowariker, N.V. Viswanathan and J. Sreedhar	Wiley-Eastern
3	Functional Monomers and Polymers	K. Takemoto, Y. Inaki and RM. Ottanbrite	
4	Contemporary Polymer Chemistry	H.R. Alcock and F.W. Lambe	Prentice Hall
5	Physics and Chemistry of Polymers	J.M.G. Cowie	J.M.G. Cowie

Semester	**IV**
Course Code	**CHM616**
Course Title	**Chemistry of Materials**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	Main aim to Study the polymers, ceramics, solid state ceramics and ionic conductors
Course Outcomes	The students will be able to: 1. Attain basic knowledge of polymers, ceramics, solid state ceramics and ionic conductors. 2. Analyze the methodologies for fabrication and characterization of nanomaterials, glasses and composites. 3. Interpretation of reaction of organic materials and non-linear materials.

UNIT I:

Introduction : Materials and their classification, Role of Chemistry in Material design.

Synthesis and characterization of materials : Preparative techniques: Ceramic methods; chemical strategies, chemical vapour deposition; preparation of nanomaterials, Langmuir-Blodgett Films. Fabrication of ordered nanostructures . Composition and purity of materials.

Multiphase Materials: Ferrous alloys; Fe-C phase transformations in ferrous alloys; stainless steels, non-ferrous alloys, properties of ferrous and non-ferrous alloys and their applications.

UNIT II:

Glasses, Ceramics, Composites and Nanomaterials: Glassy state, glass formers and glass modifiers, mechanical properties, clay products. Refractories, applications.

Microscopic composites; dispersion-strengthened and particle-reinforced, fibre-reinforced composites, macroscopic composites. Nanocrystalline phase, preparation procedures, special properties, applications, ceramic structures, characterizations, properties.

Organic Materials : Conducting organics - Metals from molecules, charge transfer materials and conducting polymers. Organic superconductors. Fullerenes. Molecular ferromagnets and ferroelectrics. Liquid crystals: mesomorphic behaviour, optical properties of liquid crystals, display devices.

Non-linear materials: Second and third order non-linear effects; molecular rectifiers and frequency doublers; unimolecular electronic devices. Photochromic materials; optical data storage, memory and switches.

UNIT III:

Polymeric Materials: Molecular shape, structure and configuration, crystallinity, stress-strain behaviour, thermal behaviour, polymer types and their applications, conducting and ferro-electric polymers.

Ionic Conductors: Types of ionic conductors, mechanism of ionic conduction, interstitial jumps. (Frenkel); vacancy mechanism, diffusion superionic conductors; phase transitions and mechanism of conduction in superionic conductors, examples and applications of ionic conductors.

UNIT IV:

High T_c Materials: Defect perovskites, high Tc superconductivity in cuprates, preparation and. characterization of 1-2-3 and 2-1-4 materials, normal state properties; anisotropy; temperature dependence of electrical resistance; optical phonon modes, superconducting state; heat capacity; coherence length, elastic constants, position lifetimes, microwave absorption-pairing and multigap structure in high Tc materials, applications of high Tc materials.

Materials for Solid State Devices: Rectifiers, transistors, capacitors -IV-V compounds, low-dimensional quantum structures, optical properties. .

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Solid State Physics	N.W. Ashcroft and N.D. Mermin,	Saunders College
2	Material Science and Engineering	An Introduction, W.D. Callister,	Wiley
3	Principles of the Solid State	H.V. Keer,	Wiley Eastern
4	Materials Science	J.C. Anderson, K.D. Leaver, J.M. Alexander and A.D. Rawlings	ELBS
5	Thermotropic Liquid Crystals	Ed., G.W. Gray	John Wiley

Semester	**IV**
Course Code	**CHM618**
Course Title	**(Photophysical Chemistry)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To introduce the concepts and importance of photochemistry and photophysical principles, their applications on simple and macromolecules.
Course Outcomes	The students will be able to: 1. assess photochemistry and photophysical principles. 2. identify and characterize of transient intermediates by ultrafast modern techniques. 3. know the theory and application of photochemistry and photophysical principles of Macromolecules

Syllabus

Unit-I

Principles and Concepts: Laws of photochemistry, Atomic and molecular term symbols, Electronic transitions, Jablonski diagram and photophysical processes, Radiative transitions, Absorption and emission, Absorption coefficient, Phosphorescence, Intersystem crossing, Mechanisms of singlet-triplet conversion (spin-orbit coupling), Spin rephasing, Spin flip.

Unit-II

Examples of ISC between states of different configurations, Radiative rates, Radiationless transitions, Internal conversion, Energy gap law, Deuterium effect. Electronically Excited States: Electronic, Vibrational and spin configurations, Excited state lifetime, Steady state and time resolved emission, Factors affecting excited state energy, Solvent effect, TICT, Origin of energy difference between singlet and triplet states, Excited state kinetics, Quantum yield.

Unit-III

Excimer and exciplex, Kinetics of luminescence quenching, Static and dynamic, Stern- Volmer analysis, Deviation from Stern-Volmer kinetics, Photoinduced electron transfer rates, Free energy dependence of electron transfer on rate, Photoinduced energy transfer, FRET, ESPIT, TBET, Rate and efficiency alculation of FRET

Unit-IV

Applications of Photochemistry and Photophyscial Principles: Measurement of fluorescence and phosphorescence and lifetimes, Introduction to time-resolved techniques for absorption and emission measurements, Detection and kinetics of reactive intermediates, Photochromic reactions and memory devices, Sensors, Switches and molecular machines, TiO_2 photocatalysis, , Flash photolysis.

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Principles of Fluorescence Spectroscopy	Lakowicz, J. R.	Springer, New York (2006), 3rd ed.
2	Fundamentals of Photoinduced Electron Transfer	Kavarnos, G. J	VCH publishers Inc., NewYork (1993).
3	Molecular Fluorescence: Principles and Applications	Valeur, B	Wiley-VCH Verlag GmbH,

			Weinheim (2002).
4	Modern Molecular Photochemistry of Organic Molecules	Turro, N. J., Ramamurthy, V., and Scaiano, J. C.,	University Science, Books, CA (2010)
5	Ninomiya, I., and Naito, T.,	Photochemical Synthesis,	Academic Press, New York (1989).

Semester	**IV**
Course Code	**CHM618**
Course Title	**(Biofuels)**
Type of course	Theory
L T P	4 0 0
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	To acquire knowledge of different methods of biofuel production, application, and their advantages.
Course Outcomes	The students will be able to: 1. know current processes for biofuel production from biomass 2. discuss the models of biomass concentration and utilization 3. know the various application of biofuels as an alternative liquid fuels.

Syllabus

Unit –I

Introduction: Drivers for alternative fuels, security, cost and environmental considerations, carbon sequestration and the impact of biofuels, review of current processes for biofuel production from biomass

Unit –II

Economic Models: Costing of current and future processes for biofuel production from biomass, biomass availability, models of biomass concentration and utilization.

Unit –III

Feedstock Chemistry: Chemistry of triglycerides and carbohydrates, Improving biomass yield and properties for easier processing and conversion, Pretreatment of biomass, Enzymatic hydrolysis, Processes and alternatives, Enzymes immobilization techniques.

Unit –IV

Fermentation: Processes and alternatives, Aqueous processing of sugars.

Bio-Diesel and other alternative liquid fuels, Policy of biofuels, Biofuels around the world: Brazil, India and China

Text and Reference books:

S.No.	Name/Title	Author	Publisher
1	Biofuels: Towards a greener and secure energy future,	*Bhojvaid, P.K.,*	TERI Press (2006).
2	Dadhich Production and Technology of Bio-diesel: Seeding a change,	Adholeya, A., and Kumar P.,	TERI press (2008).
3	Biofuels: Production, Application and Development	Scragg, A. H.,	CABI (2009).
4	Biofuels,	Olsson, L.,	Springer, (2007).
5	Biofuels: Illusion Or Reality?	Furfari, A.,	the European Experience, Editions TECHNIP (2008).

Course Code	**BOT001**
Course Title	**Natural Hazards and Disaster Management**

Type of course	Theory Course
L T P	3 0 0
Credits	3
Course prerequisite	Graduation
Course Objective	To learn about natural hazards, risk assessment and disaster management
Course Outcomes	The students will be able to: 1. know the current overview of natural hazard materials 2. discuss the physical aspects of vulnerability and elements of risk mapping, assessment 3. know the development planning, sustainable development in the context of Climate Change

Syllabus

Unit I

Overview of natural hazards; Introduction to natural hazards, impact and mitigation in Global and Indian context; causes and consequences of geological hazards, flood, drought and climate change issues, forest hazard, tsunami and coastal hazards, cyclone hazards, snow avalanche, GLOF and glacier related hazards, extreme weather events, urban and industrial hazards.

Unit II

Introduction to vulnerability and risk assessment, socio-economic and physical aspects of vulnerability and elements of risk mapping, assessment, and reduction strategies.

Unit III

Earth observation: Data availability and key operational issues for DM: EO systems for natural hazards study: present (operational) and future systems; multi-temporal data sources, multi-temporal database organisation: Key operational issues, utilisation of geo-information products for disaster management (available through International cooperation e.g. International Charter etc.)

Unit IV

Disaster management framework of India and recent initiatives by Govt. of India with special emphasis on DRR HFA 2005-2015, MDG and SAARC comprehensive framework for DRR Disaster Management Support (DMS): Status in India for use of space inputs Mainstreaming DRR in Development Planning Sustainable development in the context of Climate Change Disaster Recovery-Strategy and case examples.

Text and Reference books:

S. No.	Name/Title	Author	Publisher
1	Environmental Hazards : Assessing Risk and Reducing Disaster	Keith Smith and Petley David, 2008.	Routledge
2	Geo-information for Disaster Management	van Oosterom Peter, Zlatanova Siyka and Fendel Elfriede, 2005	Springer-Verlag
3	Geospatial Techniques in Urban Hazards and Disaster Analysis	Showalter, Pamela S. and Lu, Yongmei, 2010.	John Wiley and Sons.
4	An International Perspective on Natural Disaster: Occurrence, Mitigation and Consequences	Stoltman JP, Lidstone J and Dechano LM., 2004.	Kluwer Academic Publishers

Course Code	**CHM630**
Course Title	**Project (Phase-II)**
Type of course	Practical Course
L T P	0 0 8
Credits	4
Course prerequisite	B.Sc. with Chemistry as main subject
Course Objective	The project would develop scientific aptitude, reviewing of literature, critical thinking, hypothesis development, experiment planning, synopsis writing, problem presentation and way to solve the problem.
Course Outcomes	The students will be able to: 1. Explore research apptitude & practicalabilty of knowledge gained by student in understaning the basics of Reserach 2. Develop critical thinking through the detailed review of litrature comprehend expertise for writing the research reportsin form of review article as well as research publications. 3. Analyze & generate experimental skills towards the industrial applications. 4. Equiped for the industrial outreach through the experimental knowledge gained through project work.

- Project supervisor would be allocated at the start of the semester and research project would be undertaken in discussion with the project supervisor.
- At the end of the semester the student has to prepare a project report as per the university guidelines/ Format.
- Upon submission of the project report, the projects would be evaluated based on a project presentation.